

TOPIC – NCERT QUESTION

- Q.1** You have two solutions A and B. The pH of solution A is 6 and pH of solution B is 8. Which solution has more hydrogen ion concentration? Which of this is acidic and which one is basic?
- Q.2** What effect does the concentration of $H^+(aq)$ ions have on the nature of the solution?
- Q.3** Do basic solutions also have $H^+(aq)$ ions? If yes, then why are these basic ?

Sol.1 A pH value of less than 7 indicates an acidic solution, while greater than 7 indicates a basic solution. Since solution A has more hydrogen ion concentration, solution A is acidic and solution B is basic.

Sol.2 More the concentration of H^+ ions, higher the acidic nature of the solution.

Sol.3 Basic solutions have $H^+(aq)$ ions. But these are far less in number than OH^- ions that is responsible for their basic nature.