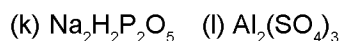
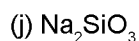
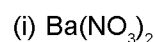
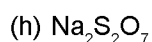
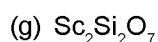
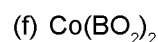
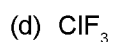
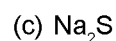
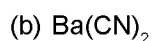
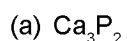


Topic : Basic Inorganic Chemistry**Type of Questions****M.M., Min.****Subjective Questions ('-1' negative marking) Q.1 to Q.4****(4 marks, 5 min.)****[16, 20]****1. Write the name of the following :****2. Write down the formula of the following :**

(a) Sulphur hexafluoride

(b) Lithium nitride

(c) strontium chloride

(d) dioxygen di fluoride

(e) barium azide

(f) barium perchlorate

(g) sodium hypochlorite

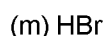
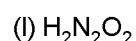
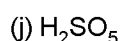
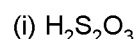
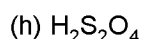
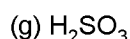
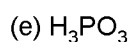
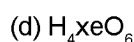
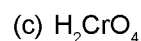
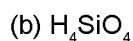
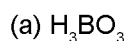
(h) calcium phosphate

(i) magnesium pyro phosphate

(j) Copper (II) metaborate

(k) Sodium pyrosulphite

(l) Ferric nitrate

3. Write the name of the following :**4. Write down the formula of the following :**

(a) Carbonic acid

(b) Pyro silicic acid

(c) Meta boric acid

(d) Manganic acid

(e) Xenic acid

(f) Hypophosphorus acid

(g) Phosphoric acid

(h) Pyrosulphurous acid

(i) Dithionic acid

(j) Chlorous acid

(k) Nitrous acid

(l) Peroxy nitric acid

(m) Hydroiodic acid

(n) Hydrocyanic acid

Answer Key

DPP No. # 6

- (a) Calcium phosphide (b) Barium cyanide (c) Sodium sulphide
 (d) Chlorine trifluoride (e) Sulphur tetrafluoride (f) Cobalt (II) metaborate
 (g) Scandium pyrosilicate (h) Sodium pyrosulphate (i) Barium nitrate
 (j) Sodium metasilicate (k) Sodium pyrophosphite (l) Aluminium sulphate
- (a) SF_6 (b) Li_3N (c) SrCl_2 (d) O_2F_2 (e) $\text{Ba}(\text{N}_3)_2$ (f) $\text{Ba}(\text{ClO}_4)_2$
 (g) NaOCl (h) $\text{Ca}_3(\text{PO}_4)_2$ (i) $\text{Mg}_2\text{P}_2\text{O}_7$ (j) $\text{Cu}(\text{BO}_2)_2$ (k) $\text{Na}_2\text{S}_2\text{O}_5$ (l) $\text{Fe}(\text{NO}_3)_3$
- (a) Ortho boric acid (b) Ortho silicic acid (c) Chromic acid
 (d) Perxenic acid (e) Phosphorus acid (f) Metaphosphoric acid
 (g) Sulphurous acid (h) Dithionous acid (i) Thiosulphuric acid
 (j) Peroxymonosulphuric acid (k) Chloric acid (l) Hyponitrous acid
 (m) Hydrobromic acid (n) Hydrazoic acid
- (a) H_2CO_3 (b) $\text{H}_6\text{Si}_2\text{O}_7$ (c) HBO_2 (d) H_2MnO_4 (e) H_2XeO_4 (f) H_3PO_2
 (g) H_3PO_4 (h) $\text{H}_2\text{S}_2\text{O}_5$ (i) $\text{H}_2\text{S}_2\text{O}_6$ (j) HClO_2 (k) HNO_2 (l) HNO_4
 (m) HI (n) HCN

Hints & Solutions

DPP No. # 6

- (a) Calcium phosphide (b) Barium cyanide (c) Sodium sulphide
 (d) Chlorine trifluoride (e) Sulphur tetrafluoride (f) Cobalt (II) metaborate
 (g) Scandium pyrosilicate (h) Sodium pyrosulphate (i) Barium nitrate
 (j) Sodium metasilicate (k) Sodium pyrophosphite (l) Aluminium sulphate
- (a) SF_6 (b) Li_3N (c) SrCl_2 (d) O_2F_2 (e) $\text{Ba}(\text{N}_3)_2$ (f) $\text{Ba}(\text{ClO}_4)_2$
 (g) NaOCl (h) $\text{Ca}_3(\text{PO}_4)_2$ (i) $\text{Mg}_2\text{P}_2\text{O}_7$ (j) $\text{Cu}(\text{BO}_2)_2$ (k) $\text{Na}_2\text{S}_2\text{O}_5$ (l) $\text{Fe}(\text{NO}_3)_3$
- (a) Ortho boric acid (b) Ortho silicic acid (c) Chromic acid
 (d) Perxenic acid (e) Phosphorus acid (f) Metaphosphoric acid
 (g) Sulphurous acid (h) Dithionous acid (i) Thiosulphuric acid
 (j) Peroxymonosulphuric acid (k) Chloric acid (l) Hyponitrous acid
 (m) Hydrobromic acid (n) Hydrazoic acid
- (a) H_2CO_3 (b) $\text{H}_6\text{Si}_2\text{O}_7$ (c) HBO_2 (d) H_2MnO_4 (e) H_2XeO_4 (f) H_3PO_2
 (g) H_3PO_4 (h) $\text{H}_2\text{S}_2\text{O}_5$ (i) $\text{H}_2\text{S}_2\text{O}_6$ (j) HClO_2 (k) HNO_2 (l) HNO_4
 (m) HI (n) HCN