

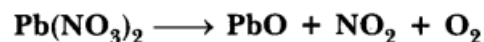
CHEMICAL REACTIONS AND EQUATIONS

VERY SHORT ANSWER TYPE QUESTIONS [1 MARK]

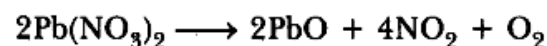
1. In electrolysis of water, why is the volume of gas collected over one electrode double that of gas collected over the other electrode?

Answer. It is because water contains hydrogen and oxygen in the ratio of 2 : 1.

2. Balance the following chemical equations.

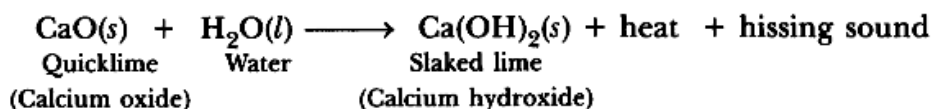


Answer.



3. What happens chemically when quicklime is added to water filled in a bucket?

Answer. Quicklime reacts with water to form slaked lime and produces lot of heat and hissing sound.



4. On what basis is a chemical equation balanced?

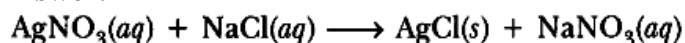
Answer. A chemical reaction is balanced on the basis of law of conservation of mass.

5. What change in colour is observed when white silver chloride is left exposed to sunlight? State the type of chemical reaction in this change.

Answer. Silver chloride becomes grey. It is a photochemical decomposition reaction.

6. Write a balanced chemical equation for the reaction between sodium chloride and silver nitrate indicating the physical state of the reactants and the products.

Answer.



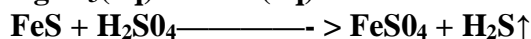
7. State one basic difference between a physical change and a chemical change.

Answer. In physical change, no new substance is formed, whereas in a chemical change, new substance(s) is/are formed.

8. What is meant by a chemical reaction?

Answer. The reaction representing a chemical change is called a chemical reaction.

9. $\text{AgNO}_3(aq) + \text{NaCl}(aq) \longrightarrow \text{AgCl}(s) \downarrow + \text{NaNO}_3(aq)$



Consider the above mentioned two chemical equations with two different kinds of arrows (\uparrow and \downarrow) along with product. What do these two different arrows indicate?

Answer. \uparrow shows the gas is evolved whereas \downarrow shows insoluble substance (precipitate) is formed.

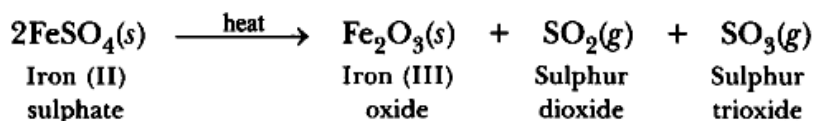
10. Hydrogen being a highly inflammable gas and oxygen being a supporter of combustion, yet water which is a compound made up of hydrogen and oxygen is used to extinguish fire. Why?

Answer. It is because properties of compound (H_2O) are different from properties of its constituting elements, i.e. H_2 and O_2 .

SHORT ANSWER TYPE QUESTIONS [I] [2 MARKS]

11. Name the products formed on strongly heating ferrous sulphate crystals. What type of chemical reaction occurs in this change?

Answer.

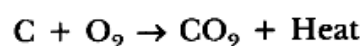
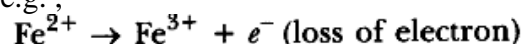


It is decomposition reaction.

12. What is an oxidation reaction? Give an example of oxidation reaction. Is oxidation an exothermic or an endothermic reaction?

Answer. The reaction in which oxygen or electronegative element is added or hydrogen or electropositive element is removed or loss of electrons takes place, is called an oxidation reaction,

e.g.,



Oxidation reactions are mostly exothermic in nature because heat is evolved in this process.

13. Describe an activity to demonstrate the change that takes place when white silver chloride is kept in sunlight. State the type of chemical reaction which takes place.

Answer.

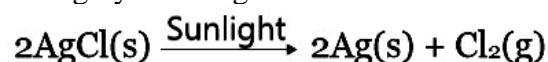
Aim: To demonstrate the change that takes place when white silver chloride is kept in sunlight.

Materials Required: $\text{AgNO}_3(\text{aq})$, $\text{NaCl}(\text{aq})$, test tubes.



Procedure:

1. Take 5 ml of silver nitrate solution in a test tube.
2. Prepare sodium chloride solution in another test tube.
3. Add sodium chloride solution into test tube containing silver nitrate solution.
4. Observe the colour of silver chloride formed. Dry it with the help of filter papers and place it on the watch glass.
5. Place the watch glass under sunlight for sometime.
6. Observe the colour of the silver chloride after sometime. Observation: White silver chloride turns grey in sunlight because silver metal is formed.

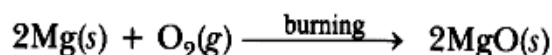


Explanation: Silver chloride is photosensitive. It decomposes in presence of sunlight to form silver metal and chlorine gas.

Conclusion: Decomposition of silver chloride in presence of sunlight is photochemical decomposition reaction.

14. When magnesium ribbon burns in air or oxygen, a product is formed. State the type of chemical reaction and name the product formed in the reaction. Write balanced chemical equation of this reaction.

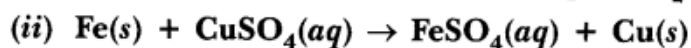
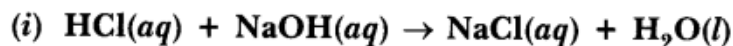
Answer.



The type of reaction is combination reaction and the product formed is magnesium oxide.

15. Distinguish between a displacement reaction and a double displacement reaction. Identify the displacement and the double displacement reaction from the following reactions.

Answer.



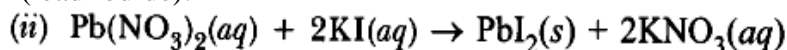
Displacement reaction is a reaction in which more reactive metal can displace less reactive metal from its salt solution.

Double displacement reaction are those reactions in which compounds exchange their ions to form two new compounds (i) Double displacement reaction (ii) Displacement reaction

16. When you have mixed the solutions of lead(II) nitrate and potassium iodide,
(i) what was the colour of the precipitate formed and can you name the precipitate?
(ii) write the balanced chemical equation for this reaction.
(iii) is this also a double displacement reaction?

Answer.

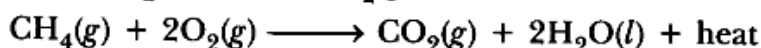
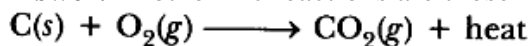
(i) The colour of the precipitate is yellow. The name of compound formed as a precipitate is PbI_2 (lead iodide).



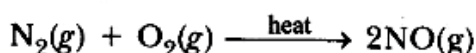
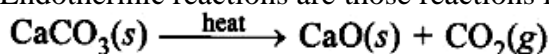
(iii) Yes, it is also a double displacement reaction.

17. What do you mean by exothermic and endothermic reactions? Give examples.

Answer. Exothermic reactions are those in which heat is evolved, e.g.



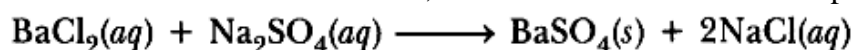
Endothermic reactions are those reactions in which heat is absorbed, e.g.



18. What happens when an aqueous solution of sodium sulphate reacts with an aqueous solution of barium chloride? State the physical conditions of reactants in which the reaction between them will not take place. Write the balanced chemical equation for the reaction and name the type of reaction.

Answer. White precipitate of barium sulphate is formed.

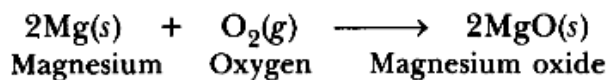
If both reactants are in solid state, then the reaction will not take place between them.



It is a double displacement as well as a precipitation reaction.

19. What is a redox reaction? When a magnesium ribbon burns in air with a dazzling flame and forms a white ash, is magnesium oxidised or reduced? Why?

Answer. The reactions in which oxidation (loss of electrons) and reduction (gain of electrons) take place simultaneously are called redox reactions.



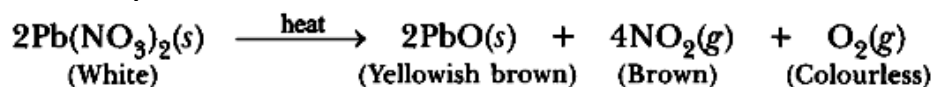
Magnesium is getting oxidised because it is losing electrons to form Mg^{2+} and oxygen is gaining electrons to form O^{2-} , therefore it is getting reduced.

20. Write any two observations in an activity which may suggest that a chemical reaction has taken place. Give an example in support of your answer.

Answer. Any two of these observations will suggest chemical reaction has taken place.

- (i) Change in state.
- (ii) Change in colour.
- (iii) Evolution of gas.
- (iv) Change in temperature.

For example, lead nitrate is white crystalline solid which on heating gives yellowish brown solid (lead monoxide). A brown gas and a colourless gas is also evolved. It shows chemical reaction has taken place.

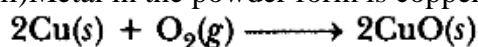


21. When the powder of a common metal is heated in an open china dish, its colour turns black. However, when hydrogen is passed over the hot black substance so formed, it regains its original colour. Based on the above information, answer the following questions.

- (i) What type of chemical reaction takes place in each of the two given steps?
- (ii) Name the metal initially taken in the powder form. Write balanced chemical equations for both reactions.

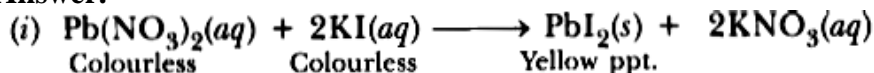
Answer.

- (i) In first step, oxidation takes place. In second step, redox reaction takes place.
- (ii) Metal in the powder form is copper.



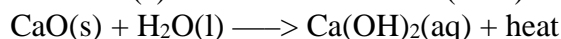
22. Using a suitable chemical equation, justify that some chemical reactions are determined by:
(i) change in colour, (ii) change in temperature.

Answer.



23. (a) A solution of substance 'X' is used for white washing. What is the substance 'X'? State the chemical reaction of 'X' with water.
(b) Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

Answer. (a) 'X' is calcium oxide (CaO).



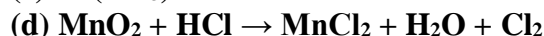
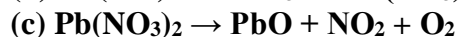
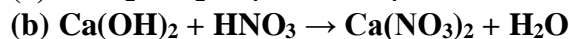
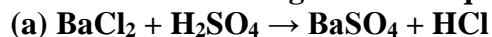
(a) It is because iron displaces copper from CuSO_4 to form FeSO_4 which is pale green.



Blue

Pale green

24. Balance the following chemical equations.



Answer. (a) $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{HCl}$

(b) $\text{Ca}(\text{OH})_2 + 2\text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$

(c) $2\text{Pb}(\text{NO}_3)_2 \rightarrow 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$

(d) $\text{MnO}_2 + 4\text{HCl} \rightarrow \text{MnCl}_2 + 2\text{H}_2\text{O} + \text{Cl}_2$

25. Write the balanced equation for the following reaction and identify the type of reaction in each case.

(a) Potassium bromide + Barium iodide \rightarrow Potassium iodide + Barium bromide.

(b) Hydrogen (g) + Chlorine(g) \rightarrow Hydrogen chloride (g)

Answer. (a) $2\text{KBr}(\text{aq}) + \text{BaI}_2(\text{aq}) \rightarrow 2\text{KI}(\text{aq}) + \text{BaBr}_2(\text{aq})$

(b) $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2\text{HCl}(\text{g})$

26. A zinc plate was put into a solution of copper sulphate kept in a glass container. It was found that blue colour of the solution gets fader and fader with the passage of time. After few days, when zinc plate was taken out of the solution, a number of holes were observed on it. (i) State the reason for changes observed on the zinc plate.

(ii) Write the chemical equation for the reaction involved.

Answer.

(i) It is because zinc has displaced copper from CuSO_4 . Zinc metal has been used to form zinc sulphate, therefore, number of holes were observed.

(ii) $\text{Zn}(\text{s}) + \text{CuSO}_4(\text{aq}) \longrightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu}(\text{s})$
Blue Colourless

27. A white salt on heating decomposes to give brown fumes and a residue is left behind.

(i) Name the salt.

(ii) Write the equation for the decomposition reaction.

Answer.

(i) Lead nitrate is white salt.

(ii) $2\text{Pb}(\text{NO}_3)_2(\text{s}) \xrightarrow{\text{heat}} 2\text{PbO}(\text{s}) + 4\text{NO}_2(\text{g}) + \text{O}_2(\text{g})$

28. When a solution of potassium iodide is added to a solution of lead nitrate in a test tube, a reaction takes place.

(a) What type of reaction is this?

(b) Write a balanced chemical equation to represent the above reaction.

Answer.

(a) Double displacement as well as precipitation reaction.

(b) $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{KI}(\text{aq}) \longrightarrow \text{PbI}_2(\text{s}) + 2\text{KNO}_3(\text{aq})$
Yellow ppt.

29. Define combination reaction. Give one example of a combination reaction which is also exothermic.

Answer. A reaction in which two elements or compounds combine to form a single compound is called combination reaction.

$\text{CaO}(\text{s}) + \text{H}_2\text{O}(\text{l}) \longrightarrow \text{Ca}(\text{OH})_2(\text{aq}) + \text{heat}$

It is also an exothermic reaction along with a combination reaction because heat is evolved.

SHORT ANSWER TYPE QUESTIONS[II] [3 MARKS]

30. Classify the following reactions into different types.

(i) $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \longrightarrow \text{AgCl}(\text{s}) + \text{NaNO}_3(\text{aq})$

(ii) $\text{CaO}(\text{s}) + \text{H}_2\text{O}(\text{l}) \longrightarrow \text{Ca}(\text{OH})_2(\text{aq})$

(iii) $2\text{KClO}_3(\text{s}) \xrightarrow{\Delta} 2\text{KCl}(\text{aq}) + 3\text{O}_2(\text{g})$

(b) Which of the above reaction(s) is/are precipitation reaction(s)? Why is a reaction called precipitation reaction?

Answer.

- (a) (i) Precipitation reaction (Double displacement reaction)
- (ii) Combination reaction (in) Decomposition reaction
- (b) Reaction (i) is a precipitation reaction because one of the product formed is insoluble in water.

31. Write balanced equations for the following mentioning the type of reaction involved.

- (i) Aluminium + Bromine \longrightarrow Aluminium bromide**
- (ii) Calcium carbonate \longrightarrow Calcium oxide + Carbon dioxide**
- (iii) Silver chloride \longrightarrow Silver + Chlorine**

Answer.

- (i) $2\text{Al}(s) + 3\text{Br}_2(g) \longrightarrow 2\text{AlBr}_3(s)$
- (ii) $\text{CaCO}_3(s) \longrightarrow \text{CaO}(s) + \text{CO}_2(g)$
- (iii) $2\text{AgCl}(s) \xrightarrow{\text{Sunlight}} 2\text{Ag}(s) + \text{Cl}_2(g)$

32. (a) Why is respiration considered as an exothermic reaction?

(b) Define the terms oxidation and reduction.

(c) Identify the substance that is oxidised and reduced in the following reaction.



Answer. (a) It is because heat is evolved during respiration.

(b) Oxidation is a process in which O_2 is added or H_2 is removed or loss of electrons take place.

Reduction is a process in which H_2 is added or O_2 is removed or gain of electrons take place.

(c) Zn is getting oxidised, CuO is getting reduced.

33. What is meant by

(i) precipitation reaction,

(ii) exothermic reaction,

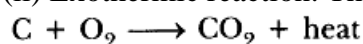
(iii) oxidation reaction?

Write balanced chemical equations for an example of each.

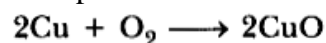
Answer.(i) Precipitation reaction: The reaction in which two compounds exchange their ions and the product formed is insoluble in water is called precipitation reaction.



(ii) Exothermic reaction: The reaction in which heat is evolved is known as exothermic reaction.



(iii) Oxidation reaction: The reaction in which O_2 is added or H_2 is removed or loss of electrons takes place is called oxidation reaction.



34. You might have noted that when copper powder is heated in a china dish, the surface of copper powder becomes coated with a black colour substance.

(i) How has this black coloured substance formed?

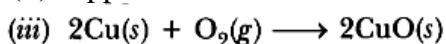
(ii) What is that black substance?

(iii) Write the chemical equation of the reaction that takes place.

Answer.

(i) Copper reacts with oxygen to form copper oxide which is black, i.e. oxidation of copper takes place.

(ii) Copper oxide

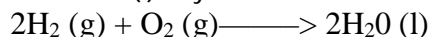


35. **“We need to balance a skeltal chemical equation.” Give reason to justify the statement.**

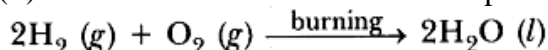
Answer. Skeltal chemical equation are unbalanced. We need to balance chemical equation because of law of conservation of mass. It states that ‘matter can neither be created nor be destroyed’. Therefore chemical equation must be balanced in each and every chemical reaction.

36. **Giving an example list two information which make a chemical equation more useful (informative).**

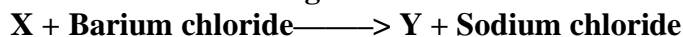
Answer. (i) Physical state of reactants must be mentioned, e.g.



(ii) Condition in which reaction takes place are written on the arrow head, e.g.



37. **Consider the following chemical reaction**

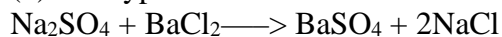


(White ppt)

(a) **Identify ‘X’ and ‘Y’** (b) **The type of reaction**

Answer. (a) ‘X’ is Na_2SO_4 and Y is BaSO_4 .

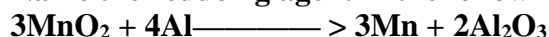
(b) The type of reaction



(White ppt)

The reaction is precipitation reaction. It is also called double displacement reaction.

38. **Name the reducing agent in the following reaction:**



State which is more reactive, Mn or Al and why?

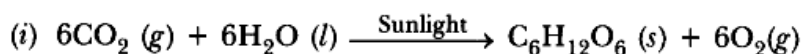
Answer. ‘Al’ is reducing agent.

‘Al’ is more reactive than Mn v ‘Al’ displaces Mn from its oxide.

39. (i) **Write a balanced chemical equation for process of photosynthesis.**

(ii) **When do desert plants take up carbon dioxide and perform photosynthesis?**

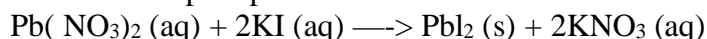
Answer.



(ii) In desert plants the stomata are open at night. They take CO_2 at night and is stored in the form of acid and is used during day time for photosynthesis.

40. **What is observed when a solution of potassium iodide solution is added to a solution of lead nitrate? Name the type of reaction. Write a balanced chemical equation to represent the above chemical reaction.**

Answer. Yellow precipitate of lead iodide is formed. It is precipitation reaction.



It is also called double displacement reaction.

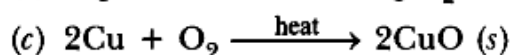
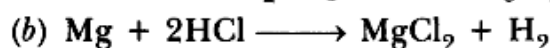
41. **Write chemical equation reactions taking place when carried out with the help of**

(a) **Iron reacts with steam**

(b) **Magnesium reacts with dil HCl**

(c) **Copper is heated in air.**

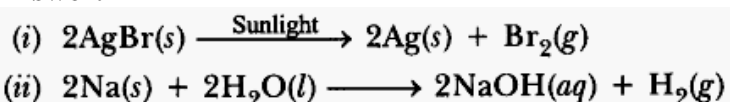
Answer.



42. Write balanced chemical equations for the following reactions.

- (i) Silver bromide on exposure to sunlight decomposes into silver and bromine,
- (ii) Sodium metal reacts with water to form sodium hydroxide and hydrogen gas.

Answer.



43. Identify the type of reaction(s) in the following equations.

- (i) $\text{CH}_4 + 2\text{O}_2 \longrightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- (ii) $\text{Pb}(\text{NO}_3)_2 + 2\text{KI} \longrightarrow \text{PbI}_2 + 2\text{KNO}_3$
- (iii) $\text{CaO} + \text{H}_2\text{O} \longrightarrow \text{Ca}(\text{OH})_2$
- (iv) $\text{CuSO}_4 + \text{Zn} \longrightarrow \text{ZnSO}_4 + \text{Cu}$

Answer.

- (i) Combustion reaction and oxidation reaction.
- (ii) Double displacement reaction and precipitation reaction.
- (iii) Combination reaction.
- (iv) Displacement reaction.

44. Write balanced equation for the reaction between magnesium and hydrochloric acid. Name the product obtained, identify the type of reaction.

Answer.



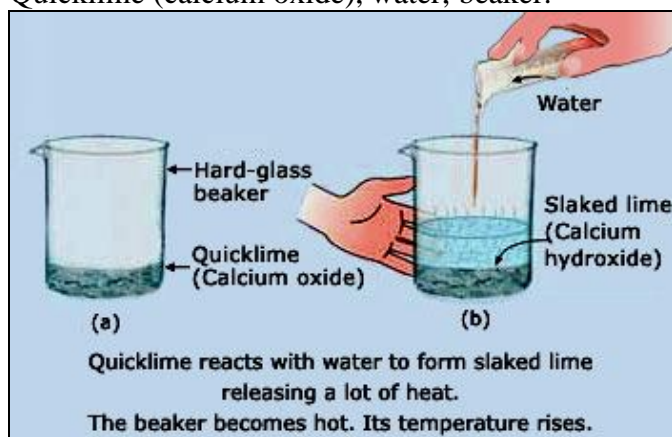
The product formed is magnesium chloride and hydrogen gas. It is a displacement reaction.

45. Describe an activity to observe what happens when quick lime is added to water taken in a beaker. State two important observations and name the type of reaction taking place.

Answer.

Aim: To observe what happens when quicklime is added to water taken in a beaker.

Materials Required:- Quicklime (calcium oxide), water, beaker.

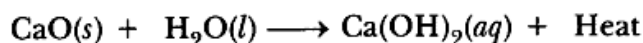


Procedure:

1. Take 5 g of calcium oxide in a beaker.
2. Add water to it slowly.
3. Touch the beaker.
4. Note down the observations.

Observation: Calcium oxide reacts with water vigorously to form calcium hydroxide with the evolution of heat.

Chemical Reaction:



Quicklime Water Slaked lime

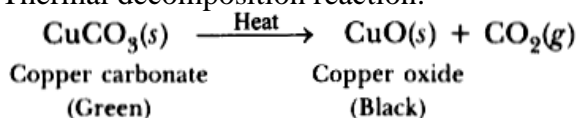
Conclusion: The reaction between CaO (Calcium oxide) and H₂O is a combination reaction. It is an exothermic process because heat is evolved.

46. What is the colour of ferrous sulphate crystals? How does this colour change after heating?

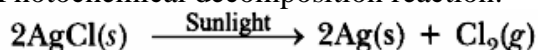
Answer. The colour of ferrous sulphate is pale green. The colour changes to reddish brown on heating due to formation of iron (III) oxide.

Give an example each for thermal decomposition and photochemical decomposition reactions. Write relevant balanced chemical equations also.

Thermal decomposition reaction:



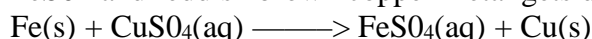
Photochemical decomposition reaction:



47. Why does the colour of copper sulphate solution change when an iron nail is dipped in it? Write two observations.

Answer. It is because displacement reaction takes place.

Iron displaces copper from copper sulphate solution and forms pale green coloured solution of FeSO₄ and reddish brown copper metal gets deposited.



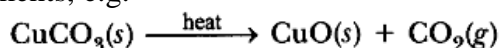
48. Translate the following statement into chemical equation and then balance it Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate. State the two types in which this reaction can be classified.

Answer. $3\text{BaCl}_2(aq) + \text{Al}_2(\text{SO}_4)_3(aq) \longrightarrow 3\text{BaSO}_4(s) + 2\text{AlCl}_3(aq)$

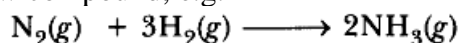
It can be classified as double displacement as well as precipitation reaction.

49. Why are decomposition reactions called the opposite of combination reactions? Write equations for these reactions.

Answer. In decomposition reaction, a compound is broken down into simpler compounds or elements, e.g.

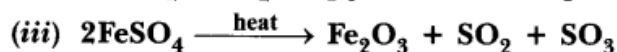
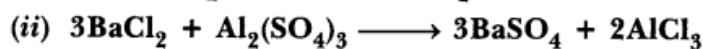
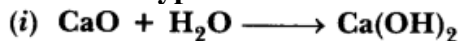


Combination reaction is a reaction in which two or more elements or compounds combine to form a new compound, e.g.



Thus, decomposition and combination reactions are opposite to each other.

50. A Name the type of chemical reaction represented by the following equation:



Answer. (i) Combination reaction

(ii) Double displacement reaction (Precipitation reaction)

(iii) Decomposition reaction.

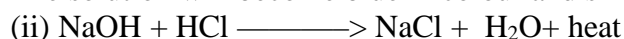
51. Write the chemical equation of the reaction in which the following changes have taken place with an example of each:

(i) Change in colour (ii) Change in temperature (iii) Formation of precipitate

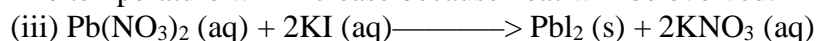
Answer.



The solution will become blue in colour and shiny silver metal will be deposited.



The temperature will increase because heat will be evolved.



Yellow ppt

Yellow precipitate of PbI_2 will be formed.

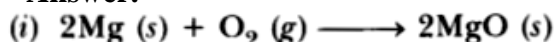
52. State the type of chemical reactions and chemical equations that take place in the following:

(i) Magnesium wire is burnt in air.

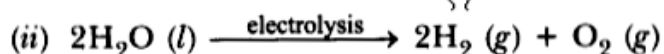
(ii) Electric current is passed through water.

(iii) Ammonia and hydrogen chloride gases are mixed.

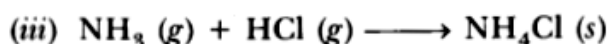
Answer.



Combination reaction (Redox reaction).

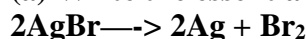


Electrical decomposition reaction.



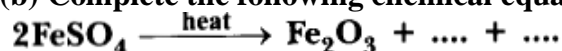
Combination reaction.

53. (a) Write the essential condition for the following reaction to take place:



Write one application of this reaction.

(b) Complete the following chemical equation of a chemical reaction

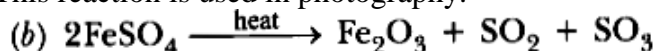


(c) What happens when water is added to quick lime. Write chemical equation.

Answer.



This reaction is used in photography.



(c) Slaked lime is formed with hissing sound and lot of heat is evolved.

54. 2g of ferrous sulphate crystals are heated in a dry boiling tube.

(i) List any two observations.

(ii) Name the type of chemical reaction taking place.

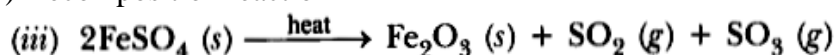
(iii) Write the chemical equation for the reaction.

Answer.

(i) *Green colour of FeSO_4 disappears and reddish brown solid is formed.

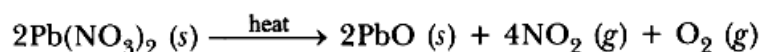
*Smell of burning sulphur.

(ii) Decomposition reaction



55. Which products will be obtained when lead nitrate is heated simply. Write balanced chemical equation for the reaction? State the type of chemical reaction that occurs in the change.

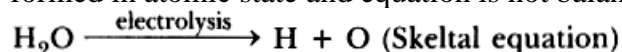
Answer. Lead monoxide, nitrogen dioxide and oxygen gas will be liberated.



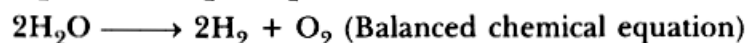
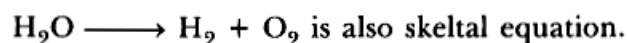
It is thermal decomposition reaction.

56. What is meant by skeltal type chemical equation? What does it represent? Using the equation for electrolytic decomposition of water, differentiate between a skeltal chemical equation and a balanced chemical equation.

Answer. The equations in which gaseous are written in atomic form instead of molecular form and equation is not balanced, are called skeltal type equation. They represent gaseous elements formed in atomic state and equation is not balanced



Hydrogen and oxygen are written in atomic forms and equation is not balanced.



57. The following diagram displays a chemical reaction. Observe carefully and answer the following questions

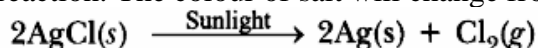


- (a) Identify the type of chemical reaction that will take place and define it. How will the colour of the salt change?

Write the chemical equation of the reaction that takes place.

- (c) Mention one commercial use of this salt.

Answer. (a) Photochemical decomposition reaction: Those reactions in which a compound breaks down into simple substances in presence of light are called photochemical decomposition reaction. The colour of salt will change from white to grey.



- (c) Silver chloride is used in photography.

58. What is rancidity? Mention any two ways by which rancidity can be prevented.

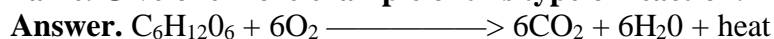
Answer. The process in which taste and smell of food gets spoiled is called rancidity. It happens due to oxidation.

Prevention from rancidity:

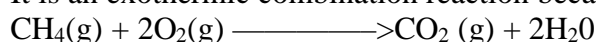
- (i) Antioxidants are added to fatty acids to prevent oxidation, e.g. chips are packed in presence of nitrogen gas which prevents spoilage by oxidation.
- (ii) Food should be kept in airtight container in refrigerator.

59. Write balanced chemical equation for the reactions that take place during respiration.

Identify the type of combination reaction that takes place during this process and justify the name. Give one more example of this type of reaction.

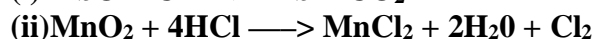
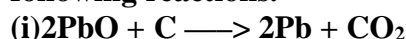


It is an exothermic combination reaction because heat is evolved.



Combustion of methane is another example of exothermic combination reaction.

60. What is redox reaction? Identify the substance oxidised and the substance reduced in the following reactions.



Answer. Those reactions in which oxidation and reduction takes place simultaneously are called redox reactions.

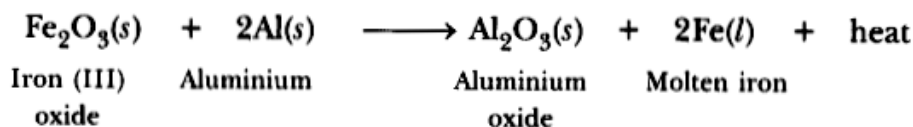
(i) PbO is getting reduced and C is getting oxidised.

(ii) MnO₂ is getting reduced and HCl is getting oxidised.

61. Write the balanced chemical equations for the following reactions and identify the type of reaction in each case.

Thermite reaction, iron (III) oxide reacts with aluminium and gives molten iron and aluminium oxide.

Answer.

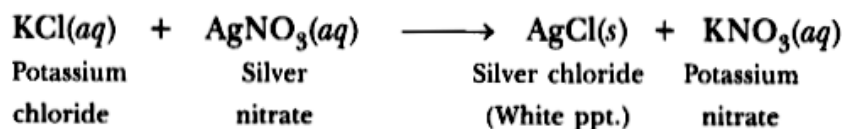


It is a displacement reaction because Al is displacing Fe from Fe₂O₃.

Molten iron is used for repairing broken railway tracks.

62. A solution of potassium chloride when mixed with silver nitrate solution, an insoluble white substance is formed. Write the chemical reaction involved and also mention the type of the chemical reaction?

Answer.



It is a double displacement reaction. It is also a precipitation reaction as AgCl is a white precipitate.

LONG ANSWER TYPE QUESTIONS [5 MARKS]

63. (a) Define a balanced chemical equation. Why should an equation be balanced?
(b) Write the balanced chemical equation for the following reaction:
(i) Phosphorus burns in presence of chlorine to form phosphorus penta chloride.
(ii) Burning of natural gas.
(iii) The process of respiration.

Answer.

(a) Balanced chemical equation has an equal number of atoms of different elements in the reactants and products. According to law of conservation of mass, matter can neither be created nor be destroyed in a chemical reaction.

(b)(i) $\text{P}_4(s) + 10\text{Cl}_2(g) \longrightarrow 4\text{PCl}_5(s)$

(i) $\text{CH}_4(g) + 2\text{O}_2(g) \longrightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l) + \text{heat energy}$

(iii) $\text{C}_6\text{H}_{12}\text{O}_6(s) + 6\text{O}_2(g) \longrightarrow 6\text{CO}_2(aq) + 12\text{H}_2\text{O}(l) + \text{energy}$

64. (a) Explain two ways by which food industries prevent rancidity.
(b) Discuss the importance of decomposition reaction in metal industry with three points.

Answer.

(a) (i) Rancidity can be prevented by adding antioxidants to food containing fat and oil, e.g. butylated hydroxy anisole is added to butter as antioxidant.

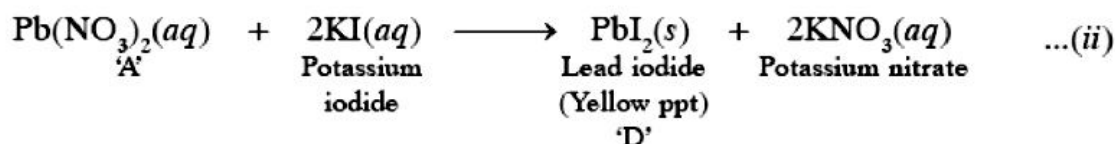
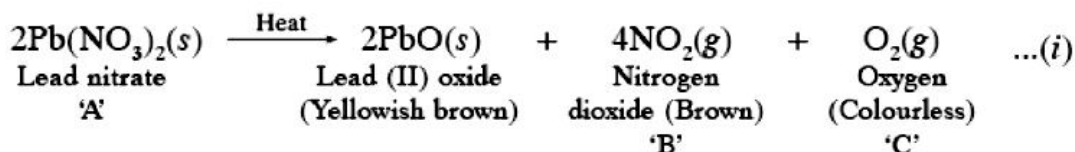
(ii) It can be prevented by packaging fat and oil containing foods in nitrogen gas.

(b) (i) Molten NaCl is electrolytically decomposed to form sodium metal.

(ii) Aluminium metal is obtained by electric decomposition of bauxite ore mixed with cryolite.

(iii) Carbonate ores are thermally decomposed to give metal oxide which on reduction give metal.

65. A metal nitrate 'A' on heating gives yellowish brown coloured metal oxide along with brown gas 'B' and a colourless gas 'C'. Aqueous solution of 'A' on reaction with potassium iodide forms a yellow precipitate of compound 'D'. Identify 'A, B, C, D'. Also identify the types of both the reactions. Metal present in 'A' is used in alloy which is used for soldering purposes.
Answer. Metal nitrate 'A' is $\text{Pb}(\text{NO}_3)_2$.

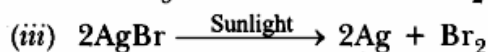
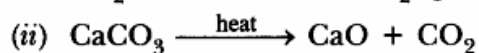
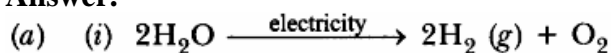


'A' is lead nitrate, 'B' is nitrogen dioxide, 'C' is oxygen and 'D' is lead iodide.

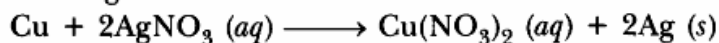
(i) is decomposition reaction and (ii) is double displacement reaction (Precipitation reaction)

66. (a) Write one example for each of decomposition reaction carried out with help of
 (i) Electricity (ii) Heat (iii) Light
 (b) Which of the following statements is correct and why copper can displace silver from silver nitrate and silver can displace copper from copper sulphate solution.

Answer.



- (b) Copper can displace silver from AgNO_3 because copper is more reactive than Ag



MULTIPLE CHOICE QUESTIONS [1 Mark]

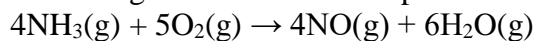
- $2\text{PbO}(s) + \text{C}(s) \rightarrow 2\text{Pb}(s) + \text{CO}_2(g)$
 Which of the following statements are correct for the above?
 a) Lead is reduced. b) Carbon dioxide is oxidized.
 c) Carbon is oxidized. d) Lead oxide is reduced.
 i) (a) and (b) ii) (a) and (c) iii) (a), (b), and (c) d) all.
- Balance the following chemical equations including the physical states.
 a) $\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow \text{C}_2\text{H}_5\text{OH} + \text{CO}_2$
 b) $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 c) $\text{NH}_3 + \text{Cl}_2 \rightarrow \text{N}_2\text{H}_4 + \text{NH}_4\text{Cl}$
 d) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
- Balance the chemical equation by including the physical states of the substances for the following reactions.
 a) Barium chloride and sodium sulphate aqueous solutions react to give insoluble Barium sulphate and aqueous solution of sodium chloride.
 b) Sodium hydroxide reacts with hydrochloric acid to produce sodium chloride and water.

c) Zinc pieces react with dilute hydrochloric acid to liberate hydrogen gas and forms zinc chloride

4. Which of the following is not a physical change?

- (a) Boiling of water to give water vapour
- (b) Melting of ice to give water
- (c) Dissolution of salt in water
- (d) Combustion of Liquefied Petroleum Gas (LPG)

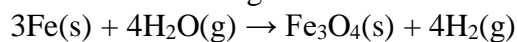
5. The following reaction is an example of a



- (i) displacement reaction
- (ii) combination reaction
- (iii) redox reaction
- (iv) neutralisation reaction

- (a) (i) and (iv) (b) (ii) and (iii) (c) (i) and (iii) (d) (iii) and (iv)

6. Which of the following statements about the given reaction are correct?



- (i) Iron metal is getting oxidised
- (ii) Water is getting reduced
- (iii) Water is acting as reducing agent
- (iv) Water is acting as oxidising agent

- (a) (i), (ii) and (iii) (b) (iii) and (iv) (c) (i), (ii) and (iv) (d) (ii) and (iv)

7. Which of the following are exothermic processes?

- (i) Reaction of water with quick lime
- (ii) Dilution of an acid
- (iii) Evaporation of water
- (iv) Sublimation of camphor (crystals)

- (a) (i) and (ii) (b) (ii) and (iii) (c) (i) and (iv) (d) (iii) and (iv)

8. Three beakers labelled as A, B and C each containing 25 mL of water were taken. A small amount of NaOH, anhydrous CuSO_4 and NaCl were added to the beakers A, B and C respectively. It was observed that there was an increase in the temperature of the solutions contained in beakers A and B, whereas in case of beaker C, the temperature of the solution falls. Which one of the following statement(s) is(are) correct?

- (i) In beakers A and B, exothermic process has occurred.
- (ii) In beakers A and B, endothermic process has occurred.
- (iii) In beaker C exothermic process has occurred.
- (iv) In beaker C endothermic process has occurred.

- (a) (i) only
(b) (ii) only
(c) (i) and (iv)
(d) (ii) and (iii)

9. A dilute ferrous sulphate solution was gradually added to the beaker containing acidified permanganate solution. The light purple colour of the solution fades and finally disappears. Which of the following is the correct explanation for the observation?

- (a) KMnO_4 is an oxidising agent, it oxidises FeSO_4
- (b) FeSO_4 acts as an oxidising agent and oxidises KMnO_4
- (c) The colour disappears due to dilution; no reaction is involved
- (d) KMnO_4 is an unstable compound and decomposes in presence of FeSO_4 to a colourless compound.

10. Which among the following is(are) double displacement reaction(s)?

- (i) $\text{Pb} + \text{CuCl}_2 \rightarrow \text{PbCl}_2 + \text{Cu}$
- (ii) $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
- (iii) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- (iv) $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

- (a) (i) and (iv)
- (b) (ii) only
- (c) (i) and (ii)
- (d) (iii) and (iv)

11. Which among the following statement(s) is(are) true? Exposure of silver chloride to sunlight for a long duration turns grey due to

- (i) the formation of silver by decomposition of silver chloride
- (ii) sublimation of silver chloride
- (iii) decomposition of chlorine gas from silver chloride
- (iv) oxidation of silver chloride

- (a) (i) only
- (b) (i) and (iii)
- (c) (ii) and (iii)
- (d) (iv) only

12. Solid calcium oxide reacts vigorously with water to form calcium hydroxide accompanied by liberation of heat. This process is called slaking of lime. Calcium hydroxide dissolves in water to form its solution called lime water. Which among the following is (are) true about slaking of lime and the solution formed?

- (i) It is an endothermic reaction
- (ii) It is an exothermic reaction
- (iii) The pH of the resulting solution will be more than seven
- (iv) The pH of the resulting solution will be less than seven

- (a) (i) and (ii)
- (b) (ii) and (iii)
- (c) (i) and (iv)
- (d) (iii) and (iv)

13. Barium chloride on reacting with ammonium sulphate forms barium sulphate and ammonium chloride. Which of the following correctly represents the type of the reaction involved?

- (i) Displacement reaction
- (ii) Precipitation reaction
- (iii) Combination reaction
- (iv) Double displacement reaction

- (a) (i) only
- (b) (ii) only
- (c) (iv) only
- (d) (ii) and (iv)

14. Electrolysis of water is a decomposition reaction. The mole ratio of hydrogen and oxygen gases liberated during electrolysis of water is
- (a) 1:1
 - (b) 2:1
 - (c) 4:1
 - (d) 1:2

15. Which of the following is(are) an endothermic process(es)?

- (i) Dilution of sulphuric acid
- (ii) Sublimation of dry ice
- (iii) Condensation of water vapours
- (iv) Evaporation of water

- (a) (i) and (iii)
- (b) (ii) only
- (c) (iii) only
- (d) (ii) and (iv)

16. In the double displacement reaction between aqueous potassium iodide and aqueous lead nitrate, a yellow precipitate of lead iodide is formed. While performing the activity if lead nitrate is not available, which of the following can be used in place of lead nitrate?

- (a) Lead sulphate (insoluble)
- (b) Lead acetate
- (c) Ammonium nitrate
- (d) Potassium sulphate

17. Which of the following gases can be used for storage of fresh sample of an oil for a long time?

- (a) Carbon dioxide or oxygen
- (b) Nitrogen or oxygen
- (c) Carbon dioxide or helium
- (d) Helium or nitrogen

18. The following reaction is used for the preparation of oxygen gas in the laboratory



Which of the following statement(s) is(are) correct about the reaction?

- (a) It is a decomposition reaction and endothermic in nature
- (b) It is a combination reaction
- (c) It is a decomposition reaction and accompanied by release of heat
- (d) It is a photochemical decomposition reaction and exothermic in nature

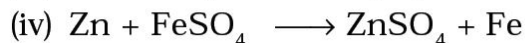
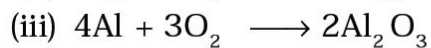
19. Which one of the following processes involve chemical reactions?

- (a) Storing of oxygen gas under pressure in a gas cylinder
- (b) Liquefaction of air
- (c) Keeping petrol in a china dish in the open
- (d) Heating copper wire in presence of air at high temperature

20. In which of the following chemical equations, the abbreviations represent the correct states of the reactants and products involved at reaction temperature?

- (a) $2\text{H}_2(\text{l}) + \text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}(\text{g})$
- (b) $2\text{H}_2(\text{g}) + \text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
- (c) $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
- (d) $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$

21. Which of the following are combination reactions?



- (a) (i) and (iii)
- (b) (iii) and (iv)
- (c) (ii) and (iv)
- (d) (ii) and (iii)

22. The removal of oxygen from a substance is called :

- (a) oxidation (b) corrosion (c) reduction (d) rancidity

23. In the context of redox reactions, the removal of hydrogen from a substance is known as :

- (a) oxidation (b) dehydration (c) reduction (d) dehydrogenation

24. The chemical reaction involved in the corrosion of iron metal is that of :

- (a) oxidation as well as displacement (b) reduction as well as combination
(c) oxidation as well as combination (d) reduction as well as displacement

25. Give the characteristic tests for the following gases

- (a) CO_2 (b) SO_2 (c) O_2 (d) H_2

.....