

CHAPTER

8

Electrochemistry

EXERCISE I (JEE MAIN)

Electrode Potential

- A metal rod is dipped in a solution of its ions. Its electrode potential is independent of
 - temperature of the solution
 - concentration of the solution
 - area of the metal exposed
 - nature of the metal
- Indicator electrode is
 - SHE
 - Calomel electrode
 - Ag/AgCl electrode
 - Quinhydrone electrode
- The position of some metals in the electrochemical series in decreasing electropositive character is given as: $\text{Mg} > \text{Al} > \text{Zn} > \text{Cu} > \text{Ag}$. What will happen if a copper spoon is used to stir a solution of aluminium nitrate?
 - The spoon will get coated with aluminium.
 - An alloy of copper and aluminium is formed.
 - The solution becomes blue.
 - No chemical change will take place.
- Four colourless salt solutions are placed in separate test tubes and a strip of copper is placed in each. Which solution finally turns blue?
 - AgNO_3
 - $\text{Pb(NO}_3)_2$
 - $\text{Zn(NO}_3)_2$
 - $\text{Cd(NO}_3)_2$
- The metal that cannot be obtained on reduction of its oxide by aluminium is
 - K
 - Mn
 - Cr
 - Fe
- Beryllium is placed above magnesium in the group II. Beryllium dust, therefore, when added to MgCl_2 solution will
 - have no effect
 - precipitate Mg metal
 - precipitate MgO
 - lead to dissolution of Be metal
- A standard hydrogen electrode has zero electrode potential because
 - hydrogen is easiest to oxidize
 - this electrode potential is assumed to be zero
 - hydrogen atom has only one electron
 - hydrogen is the lightest element

8. The standard reduction potential values of three metallic cations, X, Y and Z are +0.52, -3.03 and -1.18 V, respectively. The order of reducing power of the corresponding metals is
- (a) $Y > Z > X$
 - (b) $X > Y > Z$
 - (c) $Z > Y > X$
 - (d) $Z > X > Y$
9. A gas 'X' at 1 atm is bubbled through a solution containing a mixture of 1 M $-Y^-$ and 1 M $-Z^-$ at 25°C. If the reduction potential of $Z > Y > X$, then,
- (a) Y will oxidize X and not Z
 - (b) Y will oxidize Z and not X
 - (c) Y will oxidize both X and Z
 - (d) Y will reduce both X and Z
10. Standard electrode potential data are useful for understanding the suitability of an oxidant in a redox titration. Some half-cell reactions and their standard potentials are given below:
- $$\text{MnO}_4^- (\text{aq}) + 8\text{H}^+ (\text{aq}) + 5\text{e}^- \rightarrow \text{Mn}^{2+} (\text{aq}) + 4\text{H}_2\text{O} (\text{l}) \quad E^\circ = 1.51 \text{ V}$$
- $$\text{Cr}_2\text{O}_7^{2-} (\text{aq}) + 14\text{H}^+ (\text{aq}) + 6\text{e}^- \rightarrow 2\text{Cr}^{3+} (\text{aq}) + 7\text{H}_2\text{O} (\text{l}) \quad E^\circ = 1.38 \text{ V}$$
- $$\text{Fe}^{3+} (\text{aq}) + \text{e}^- \rightarrow \text{Fe}^{2+} (\text{aq}) \quad E^\circ = 0.77 \text{ V}$$
- $$\text{Cl}_2 (\text{g}) + 2\text{e}^- \rightarrow 2\text{Cl}^- (\text{aq}) \quad E^\circ = 1.40 \text{ V}$$
- Identify the only incorrect statement regarding the quantitative estimation of aqueous $\text{Fe}(\text{NO}_3)_2$
- (a) MnO_4^- can be used in aqueous HCl
 - (b) $\text{Cr}_2\text{O}_7^{2-}$ can be used in aqueous HCl
 - (c) MnO_4^- can be used in aqueous H_2SO_4
 - (d) $\text{Cr}_2\text{O}_7^{2-}$ can be used in aqueous H_2SO_4
11. The decreasing order of standard electrode potential of Mg, K, Ba and Ca, is
- (a) K, Ca, Ba, Mg
 - (b) Ba, Ca, K, Mg
 - (c) Ca, Mg, K, Ba
 - (d) Mg, Ca, Ba, K
12. A metal having negative reduction potential, when dipped in the solution of its own ions, has a tendency to
- (a) remain as metal atoms
 - (b) become electrically positive
 - (c) become electrically negative
 - (d) be deposited from the solution
13. The calomel electrode is reversible with respect to
- (a) Hg_2^{2+}
 - (b) H^+
 - (c) Hg^{2+}
 - (d) Cl^-
14. Which one of the following does not get oxidized by bromine water?
- (a) Fe^{2+} to Fe^{3+}
 - (b) Cu^+ to Cu^{2+}
 - (c) Mn^{2+} to MnO_4^-
 - (d) Sn^{2+} to Sn^{4+}
15. In salt bridge, normally KCl is used because
- (a) it is a strong electrolyte.
 - (b) it is good conductor of electricity.
 - (c) K^+ and Cl^- ions have nearly same ionic mobility.
 - (d) it is an ionic compound.
16. By how much would the oxidizing power of $\text{MnO}_4^-/\text{Mn}^{2+}$ couple change if the H^+ ions concentration is decreased 100 times at 25°C?
- (a) increases by 189 mV
 - (b) decreases by 189 mV
 - (c) will increase by 19 mV
 - (d) will decrease by 19 mV
17. The solution of CuSO_4 , in which copper rod is immersed, is diluted to 10 times. The reduction electrode potential
- (a) increases by 0.0295 V
 - (b) decreases by 0.0295 V
 - (c) increases by 0.059 V
 - (d) decreases by 0.059 V

18. The standard reduction potential of oxygen in acidic solution is +1.23 V. What is the standard reduction potential of oxygen in basic solution?
 (a) +0.404 V (b) -0.404 V
 (c) +2.056 V (d) -2.056 V
19. The standard reduction potentials of $\text{Cu}^{2+} | \text{Cu}$ and $\text{Cu}^{2+} | \text{Cu}^+$ are 0.337 V and 0.153 V, respectively. The standard electrode potential of $\text{Cu}^+ | \text{Cu}$ half-cell is
 (a) 0.184 V
 (b) 0.827 V
 (c) 0.521 V
 (d) 0.490 V
20. The electrode potential of hydrogen electrode in neutral solution and 298 K is
 (a) -0.413 V (b) zero
 (c) -0.826 V (d) +0.413 V
21. Electrode potential will be more for hydrogen electrode at pH (at the same temperature)
 (a) 4 (b) 3
 (c) 2 (d) 5
22. Saturated solution of KNO_3 is used to make 'salt bridge' because
 (a) velocity of K^+ is greater than that of NO_3^-
 (b) velocity of NO_3^- is greater than that of K^+
 (c) velocities of both K^+ and NO_3^- are nearly the same
 (d) KNO_3 is highly soluble in water
23. The standard reduction potentials of $\text{Pt} | \text{Cr}_2\text{O}_7^{2-}, \text{Cr}^{+3}; \text{Pt} | \text{MnO}_4^-, \text{Mn}^{+2}; \text{Pt} | \text{Ce}^{+4}, \text{Ce}^{+3}$ in the presence of acid are 1.33 V, 1.51 V and 1.61 V, respectively, at 25°C. The decreasing order of oxidizing power is
 (a) $\text{Cr}_2\text{O}_7^{2-} > \text{MnO}_4^- > \text{Ce}^{+4}$
 (b) $\text{MnO}_4^- > \text{Cr}_2\text{O}_7^{2-} > \text{Ce}^{+4}$
 (c) $\text{Ce}^{+4} > \text{MnO}_4^- > \text{Cr}_2\text{O}_7^{2-}$
 (d) $\text{MnO}_4^- > \text{Ce}^{+4} > \text{Cr}_2\text{O}_7^{2-}$
24. The standard reduction potentials at 25°C of $\text{Li}^+ | \text{Li}$, $\text{Ba}^{2+} | \text{Ba}$, $\text{Na}^+ | \text{Na}$ and $\text{Mg}^{2+} | \text{Mg}$ are -3.05, -2.73, -2.71 and -2.37 V, respectively. Which is the strongest reducing agent?
 (a) Li (b) Ba
 (c) Na (d) Mg
25. Some standard electrode potentials are given:
 $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}; E^\circ = -0.440 \text{ V}$
 $\text{Fe}^{3+} + 3\text{e}^- \rightarrow \text{Fe}; E^\circ = -0.036 \text{ V}$
 The standard electrode potential for: $\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}^{2+}$, is
 (a) -0.476 V (b) -0.404 V
 (c) +0.988 V (d) +0.772 V
26. The dissociation constant for CH_3COOH is 1.8×10^{-5} at 298 K. The electrode potential for the half-cell: $\text{Pt} | \text{H}_2 (1 \text{ bar}) | 0.5 \text{ M} - \text{CH}_3\text{COOH}$, at 298 K is ($\log 2 = 0.3$; $\log 3 = 0.48$; $2.303RT/F = 0.06$)
 (a) -0.3024 V
 (b) -0.1512 V
 (c) +0.3024 V
 (d) +0.1512 V
27. At 25°C, the solubility product of CuCl is 2.0×10^{-7} and $E^\circ_{\text{Cl}^- | \text{CuCl} | \text{Cu}}$ is 0.128 V. The value of $E^\circ_{\text{Cu}^+ | \text{Cu}}$ is ($\log 2 = 0.3$; $2.303RT/F = 0.06$)
 (a) -0.274 V (b) -0.402 V
 (c) +0.53 V (d) +0.402 V
28. The standard reduction potential of oxygen in acidic solution is 1.23 V ($\text{O}_2 + 4\text{H}_3\text{O}^+ + 4\text{e}^- \rightarrow 6\text{H}_2\text{O}$). The standard reduction potential of oxygen in basic solution is ($2.303RT/F = 0.06$)
 (a) -1.23 V
 (b) -0.39 V
 (c) +0.39 V
 (d) +2.07 V

29. The standard potentials of $\text{MnO}_4^- | \text{Mn}^{2+}$ and $\text{MnO}_2 | \text{Mn}^{2+}$ electrodes in acid solution are 1.51 and 1.23 V, respectively. Standard electrode potential for the electrode, $\text{MnO}_4^- | \text{MnO}_2$ in acid solution is
- +1.697 V
 - +5.09 V
 - +0.28 V
 - +1.37 V
30. The following reactions represent the reduction of IO_3^- ion into I^- ion in acidic and basic medium. Predict in which medium IO_3^- ion will act as a better oxidizing agent?
- $$\text{IO}_3^- + 6\text{H}^+ + 6\text{e}^- \rightarrow \text{I}^- + 3\text{H}_2\text{O}; E^\circ = +0.907 \text{ V}$$
- $$\text{IO}_3^- + 3\text{H}_2\text{O} + 6\text{e}^- \rightarrow \text{I}^- + 6\text{OH}^-; E^\circ = +0.260 \text{ V}$$
- Acid medium
 - Basic medium
 - Equally in both
 - Not predictable

Galvanic Cell

31. The correct cell diagram for the following reaction and E° for the cell is
- $$2\text{AgBr(s)} + \text{H}_2(\text{g}) \rightarrow 2\text{Ag(s)} + 2\text{H}^+ + 2\text{Br}^-$$
- $$E^\circ_{\text{AgBr}|\text{Ag}|\text{Br}^-} = +0.10 \text{ V.}$$
- (Pt) $\text{H}_2 | \text{H}^+ || \text{Br}^- | \text{AgBr} | \text{Br}_2$ (Pt); $E^\circ = 0.10 \text{ V}$
 - (Pt) $\text{H}_2 | \text{H}^+ || \text{Br}^- | \text{AgBr} | \text{Br}_2$ (Pt); $E^\circ = -0.10 \text{ V}$
 - (Pt) $\text{Br}_2 | \text{AgBr} | \text{Br}^- || \text{H}^+ | \text{H}_2$ (Pt); $E^\circ = 0.10 \text{ V}$
 - (Pt) $\text{Br}_2 | \text{AgBr} | \text{Br}^- || \text{H}^+ | \text{H}_2$ (Pt); $E^\circ = -0.10 \text{ V}$
32. In an experimental set-up for the measurement of EMF of a half-cell using a reference electrode and a salt bridge, when the salt bridge is removed, the voltage
- remains the same
 - increases to maximum
 - decreases half the value
 - drops to zero
33. After some time, the voltage of an electrochemical cell becomes zero. This is because
- their electrode potential becomes zero.
 - their reduction potential become equal but have opposite sign.
 - their reduction potential become equal and have the same sign.
 - the ions of the electrolyte in the salt bridge stop moving.
34. The cell reaction for the given cell is spontaneous if Pt, Cl_2 (P_1 atm) $| \text{Cl}^- | \text{Cl}_2$ (P_2 atm), Pt
- $P_1 > P_2$
 - $P_1 < P_2$
 - $P_1 = P_2$
 - $P_1 = 1 \text{ atm}$
35. Which one of the following statements is incorrect regarding an electrochemical cell?
- The electrode on which oxidation takes place is called anode.
 - Anode is the negative pole.
 - The direction of the current is same as that of the direction of flow of electrons.
 - The flow of current is partly due to flow of electrons and partly due to flow of ions.
36. When an electric current is drawn from a galvanic cell
- EMF suddenly increases.
 - EMF gradually increases and attains a maximum value.
 - EMF decreases and finally falls to zero.
 - EMF must remain constant.

37. Identification of anode and cathode in an electrochemical cell is made by the use of
 (a) Galvanometer
 (b) Salt bridge
 (c) Voltmeter
 (d) Potentiometer
38. For the cell $\text{Zn} | \text{Zn}^{2+} || \text{Cu}^{2+} | \text{Cu}$, if the concentration of both, Zn^{2+} and Cu^{2+} ions are doubled, the EMF of the cell
 (a) doubles
 (b) reduces to half
 (c) remains same
 (d) becomes zero
39. The standard EMF of a galvanic cell can be calculated from
 (a) the size of the electrode
 (b) the pH of the solution
 (c) the amount of metal in the anode
 (d) the E° values of the half-cells
40. The value of equilibrium constant for a feasible cell reaction must be
 (a) <1
 (b) Zero
 (c) $=1$
 (d) >1
41. If the cell reaction is spontaneous then
 (a) $\Delta G^\circ = -ve$
 (b) $E^\circ_{\text{red}} = -ve$
 (c) $E^\circ_{\text{red}} = +ve$
 (d) $\Delta G = -ve$
42. Which of the following pair of metals, when coupled, will give maximum EMF for a voltaic cell?
 (a) Fe and Cu
 (b) Pb and Au
 (c) Cu and Au
 (d) Ca and Cu
43. When a lead storage battery is charged
 (a) PbO_2 dissolves.
 (b) the lead electrode becomes coated with lead sulphate.
 (c) sulphuric acid is regenerated.
 (d) the amount of acid decreases.
44. Use of lithium metal as an electrode in high energy density batteries is due to
 (a) lithium is the lightest element.
 (b) lithium has the highest oxidation potential.
 (c) lithium is quite reactive.
 (d) lithium does not corrode readily.
45. A depolarizer used in dry cell batteries is
 (a) NH_4Cl
 (b) MnO_2
 (c) KOH
 (d) Na_3PO_4
46. Which is correct about fuel cells?
 (a) Cell continuously run as long as fuels are supplied.
 (b) These are more efficient and free from pollution.
 (c) These are used to provide power and drinking water to astronauts in space programme.
 (d) All of these
47. When a lead storage battery is discharged
 (a) SO_2 is evolved
 (b) lead sulphate is consumed
 (c) lead is formed
 (d) sulphuric acid is consumed
48. For a cell reaction involving a two-electron change, the standard EMF of the cell is found to be 0.295 V at 25°C . The equilibrium constant of the reaction at 25°C will be
 (a) 1×10^{10}
 (b) 1×10^{-10}
 (c) 29.5×10^{-2}
 (d) 2×10^{10}
49. The E_{Cell} for $\text{Ag(s)} | \text{AgI (satd)} || \text{Ag}^+ (0.10 \text{ M}) | \text{Ag (s)}$ is +0.413 V. What is the value of K_{sp} of AgI?
 (a) 1.0×10^{-8}
 (b) 1.0×10^{-7}
 (c) 1.0×10^{-14}
 (d) 1.0×10^{-16}
50. Assuming that hydrogen behaves as an ideal gas, what is the EMF of the cell at 25°C if $P_1 = 600 \text{ mm}$ and $P_2 = 420 \text{ mm}$: $\text{Pt} | \text{H}_2 (P_1) | \text{HCl} | \text{H}_2 (P_2) | \text{Pt}$? [Given: $2.303RT/F = 0.06$, $\log 7 = 0.85$]
 (a) -0.0045 V
 (b) -0.0 V
 (c) $+0.0045 \text{ V}$
 (d) $+0.0015 \text{ V}$

51. For the electrochemical cell, $M | M^+ || X^- | X$, $E_{M^+|M}^\circ = 0.44 \text{ V}$ and $E_{X|X^-}^\circ = 0.33 \text{ V}$. From this data, one can deduce that
- $M + X \rightarrow M^+ + X^-$ is the spontaneous reaction
 - $M^+ + X^- \rightarrow M + X$ is the spontaneous reaction
 - $E_{\text{cell}} = 0.77 \text{ V}$
 - $E_{\text{cell}} = -0.77 \text{ V}$
52. The EMF of the cell: $Zn | Zn^{2+} (0.01 \text{ M}) || Fe^{2+} (0.001 \text{ M}) | Fe$ at 298 K is 0.2905 V , then the value of equilibrium constant for the cell reaction is
- $e^{\frac{0.32}{0.0295}}$
 - $10^{\frac{0.32}{0.0295}}$
 - $10^{\frac{0.26}{0.0295}}$
 - $10^{\frac{0.32}{0.0591}}$
53. For the cell reaction: $4Br^- + O_2 + 4H^+ \rightleftharpoons 2Br_2 + 2H_2O$; $E^\circ = 0.18 \text{ V}$. The value of $(\log K_c)$ at 298 K is $[2.303RT/F = 0.06]$
- 12
 - 6
 - 18
 - 3
54. The standard EMF for the cell reaction: $Zn(s) + Cu^{2+}(aq) \rightarrow Zn^{2+}(aq) + Cu(s)$ is 1.10 volts at 25°C . The EMF of the cell reaction when $0.1 \text{ M } Cu^{2+}$ and $0.1 \text{ M } Zn^{2+}$ solutions are used at 25°C is
- 1.10 V
 - 1.041 V
 - -1.10 V
 - -1.041 V
55. For the cell: $Ni | Ni^{2+} || Cu^{2+} | Cu$; $E^\circ = 0.77 \text{ V}$. By which of the following activity, E_{cell} will increase?
- On decreasing $[Ni^{2+}]$
 - On decreasing $[Cu^{2+}]$
 - On increasing mass of Ni electrode
 - On increasing mass of Cu electrode
56. The standard EMF of a Daniel cell at 298 K is E_1 . When the concentration of $ZnSO_4$ is 1.0 M and that of $CuSO_4$ is 0.01 M , the EMF becomes E_2 at 298 K . The correct relationship between E_1 and E_2 is
- $E_1 = E_2$
 - $E_2 = 0$
 - $E_1 > E_2$
 - $E_1 < E_2$
57. For a reaction $A(s) + 2B^+ \rightarrow A^{2+} + B(s)$; K_c has been found to be 10^{12} . The EMF of the cell is
- 0.354 V
 - 0.708 V
 - 0.534 V
 - 0.453 V
58. An electrochemical cell is set up as follows:
- $$Pt | H_2(s) (1 \text{ atm}) | 0.001 \text{ M } HCl || 0.1 \text{ M } HA | H_2(s) (1 \text{ atm}) | Pt$$
- EMF of this cell is zero because (pK_a of $HA = 5$)
- molar concentrations of acids are different
 - Temperature is constant
 - pH of two solutions are same
 - both are standard hydrogen electrodes
59. The reaction: $H_2(g) + 2AgCl(s) \rightarrow 2H^+(aq) + 2Cl^-(aq) + 2Ag(s)$ occurs in the galvanic cell
- $Ag | AgCl(s) | KCl(aq) | AgNO_3(aq) | Ag$
 - $Pt | H_2(g) | HCl(aq) | AgNO_3(aq) | Ag$
 - $Pt | H_2(g) | HCl(aq) | AgCl(s) | Ag$
 - $Pt | H_2(g) | KCl(aq) | AgCl(s) | Ag$
60. The standard reduction potentials in acidic conditions are 0.77 V and 0.53 V , respectively, for $Fe^{3+} | Fe^{2+}$ and $I_3^- | I^-$ couples. The equilibrium constant for the reaction: $2Fe^{3+} + 3I^- \rightleftharpoons 2Fe^{2+} + I_3^-$, is $(2.303RT/F = 0.06)$
- 2×10^8
 - 10^8
 - 10^4
 - 10^{-8}
61. The following electrochemical cell has been set up: $Pt(s) | Fe^{3+}, Fe^{2+} (a = 1) || Ce^{4+}, Ce^{3+} (a = 1) | Pt(s)$; $E^\circ(Fe^{3+}|Fe^{2+}) = 0.77 \text{ V}$; $E^\circ(Ce^{4+}|Ce^{3+}) = 1.61 \text{ V}$. If an ammeter is connected between the two platinum electrodes, predict the direction of flow of current. Will the current increase or decrease with time?
- Ce electrode to Fe electrode, decrease
 - Ce electrode to Fe electrode, increase
 - Fe electrode to Ce electrode, decrease
 - Fe electrode to Ce electrode, increase

62. For the reaction: H_2 (1 bar) + $2\text{AgCl(s)} \rightleftharpoons 2\text{Ag(s)} + 2\text{H}^+$ (0.1 M) + 2Cl^- (0.1 M); $\Delta G^\circ = -48,250 \text{ J}$ at 25°C . The EMF of cell in which the given reaction takes place is
 (a) 0.25 V
 (b) 0.37 V
 (c) 0.13 V
 (d) 0.049 V
63. A cell contains two hydrogen electrodes. The negative electrode is in contact with a solution of 10^{-6} M hydrogen ions. The EMF of the cell is 0.118 V at 25°C . The concentration of hydrogen ions at the positive electrode is
 (a) 10^{-6} M
 (b) 10^{-3} M
 (c) 10^{-4} M
 (d) 10^{-5} M
64. The standard potentials of OCl^-/Cl^- and Cl^-/Cl_2 are 0.94 V and -1.36 V , respectively. The E° value of OCl^-/Cl_2 will be
 (a) 3.24 V
 (b) -0.42 V
 (c) -2.30 V
 (d) 0.52 V
65. From the following E° values for the half-cells:
 (i) $\text{D} \rightarrow \text{D}^{2+} + 2\text{e}^-$; $E^\circ = -1.5 \text{ V}$
 (ii) $\text{B}^+ + \text{e}^- \rightarrow \text{B}$; $E^\circ = -0.5 \text{ V}$
 (iii) $\text{A}^{3-} \rightarrow \text{A}^{2-} + \text{e}^-$; $E^\circ = 1.5 \text{ V}$
 (iv) $\text{C}^{2+} + \text{e}^- \rightarrow \text{C}^+$; $E^\circ = +0.5 \text{ V}$
- Which combination of two half-cells would result in a cell with largest potential?
 (a) i and iii
 (b) i and iv
 (c) iii and iv
 (d) ii and iv

Electrolysis

66. Which of the following statements does not differentiate between electrochemical cell and electrolytic cell?
 (a) Spontaneous or non-spontaneous nature of the chemical process.
 (b) Chemical reactions occurring at the electrodes
 (c) Positive and negative nature of anode.
 (d) EMF measurement.
67. The electrode through which electrons enter the electrolytic solution is
 (a) cathode
 (b) anode
 (c) may be anode or cathode
 (d) both, anode and cathode
68. Which process occurs in the electrolysis of an aqueous solution of nickel chloride at nickel anode?
 (a) $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$
 (b) $\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$
 (c) $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$
 (d) $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$
69. If mercury is used as a cathode during the electrolysis of an aqueous NaCl solution, the ions discharged at cathode are
 (a) H^+
 (b) Na^+
 (c) OH^-
 (d) Cl^-
70. Electrochemical equivalent is more for
 (a) Hydrogen
 (b) Silver
 (c) Copper
 (d) Zinc
71. The electrolytic bath used in gold plating of copper articles contains
 (a) Molten gold
 (b) Copper sulphate(aq)
 (c) $\text{AuCl}_3(\text{aq})$
 (d) $\text{AuCl}_3 + \text{NaCN}(\text{aq})$

72. Copper can be deposited from acidified copper sulphate and alkaline cuprous cyanide. If the same current is passed for a definite time
- the amount of copper deposited from acidic copper sulphate will be higher
 - the amount of copper deposited from alkaline cuprous cyanide will be higher
 - the same amount of copper will be deposited
 - copper will not deposit in either case
73. In the electrolytic cell, flow of electrons is from
- Cathode to anode in solution
 - Cathode to anode through external supply
 - Cathode to anode through internal supply
 - Anode to cathode through internal supply
74. Electrolytic cell is used to convert
- Chemical energy to electrical energy
 - Electrical energy to chemical energy
 - Chemical energy to mechanical energy
 - Electrical energy to mechanical energy
75. Faraday's law of electrolysis fails when
- temperature is increased
 - inert electrodes are used
 - a mixture of electrolytes is used
 - in none of these cases
76. Using same quantity of current, which among Na, Mg and Al is deposited more (by mass) during electrolysis of their molten salts?
- Na
 - Mg
 - Al
 - All in same
77. A certain current liberated 0.50 g of hydrogen in 2 h. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphate solution? ($\text{Cu} = 63.5$)
- 12.7 g
 - 15.88 g
 - 31.75 g
 - 63.5 g
78. A current of 3.7 A is passed for 6 h between nickel electrodes in 0.50 l of 2 M solution of $\text{Ni}(\text{NO}_3)_2$. The molarity of Ni^{2+} at the end of electrolysis is
- 1.172 M
 - 0.172 M
 - 0.586 M
 - 2 M
79. The current efficiency of an electrodeposition of copper metal in which 9.8 g of copper is deposited by a current of 3 A for 10000 s, from aqueous copper sulphate solution, is about
- 60%
 - 99%
 - 92%
 - 75%
80. On passing electricity through dilute H_2SO_4 solution, the mass of substances liberated at the cathode and anode are in the ratio of
- 1:8
 - 8:1
 - 1:32
 - 1:16
81. The electrochemical equivalents of two substances are E_1 and E_2 . The current that must pass to deposit the same amount at the cathodes in the same time must be in the ratio of
- $E_1:E_2$
 - $E_2:E_1$
 - $(E_1 - E_2):E_2$
 - $E_1:(E_2 - E_1)$
82. The same quantity of electricity is passed through one molar solution of H_2SO_4 and one molar solution of HCl. The amount of hydrogen evolved from H_2SO_4 as compared to that from HCl is
- the same
 - twice as such
 - one half as such
 - dependent on size of electrode

83. In the electrolysis of acidified AgNO_3 solution using Pt-electrodes, the anode reaction is
 (a) $2\text{NO}_3^- \rightarrow 2\text{NO}_2 + \text{O}_2 + 2\text{e}^-$
 (b) $\text{NO}_3^- \rightarrow \text{NO} + \frac{1}{2}\text{O}_2 + \text{e}^-$
 (c) $2\text{H}_2\text{O} \rightarrow 4\text{H}^+ + \text{O}_2 + 4\text{e}^-$
 (d) $\text{Pt} \rightarrow \text{Pt}^{3+} + 3\text{e}^-$
84. Two platinum electrodes were immersed in a solution of CuSO_4 and electric current was passed through the solution. After some time, it was found that colour of CuSO_4 disappeared with the evolution of gas at the electrode. The colourless solution contains
 (a) platinum sulphate
 (b) copper sulphate
 (c) copper hydroxide
 (d) sulphuric acid
85. A solution containing 1.0 M each of $\text{Cu}(\text{NO}_3)_2$, $\text{Mg}(\text{NO}_3)_2$, AgNO_3 , $\text{Hg}(\text{NO}_3)_2$ is being electrolysed using inert electrodes. The values of standard electrode potential are: $\text{Ag}^+|\text{Ag} = 0.80\text{ V}$, $\text{Hg}^{2+}|\text{Hg} = 0.79\text{ V}$, $\text{Cu}^{2+}|\text{Cu} = 0.34\text{ V}$, $\text{Mg}^{2+}|\text{Mg} = -2.37\text{ V}$. With increasing voltage, the sequence of deposit of metals on the cathode will be
 (a) Ag, Hg, Cu, Mg
 (b) Mg, Cu, Hg, Ag
 (c) Ag, Mg, Cu
 (d) Cu, Hg, Ag
86. During the electrolysis of an aqueous salt solution, the pH in the space near one of the electrode was increased and the other one was decreased. The salt solution was
 (a) NaCl (very dilute)
 (b) ZnCl_2
 (c) NaCl (Conc.)
 (d) $\text{Cu}(\text{NO}_3)_2$
87. Two electrolytic cells, one containing acidified ferrous chloride and another acidified ferric chloride, are connected in series. The mass ratio of iron deposited at cathodes in the two cells will be
 (a) 3:1
 (b) 2:3
 (c) 1:1
 (d) 3:2
88. A galvanic cell is set up from a zinc bar weighing 100 g and 1.0 L of 1.0 M copper sulphate solution. How long would the cell run if it is assumed to deliver a steady current of 1.0 A? ($\text{Zn} = 65.4$)
 (a) 53.6 h
 (b) 26.8 h
 (c) 81.97 h
 (d) 40.99 h
89. An ion is reduced to the element when it absorbs 6×10^{20} electrons. The number of equivalents of the ion is
 (a) 0.10
 (b) 0.01
 (c) 0.001
 (d) 0.0001
90. In the lead storage battery, the anode reaction is $\text{Pb(s)} + \text{HSO}_4^- + \text{H}_2\text{O} \rightarrow \text{PbSO}_4(\text{s}) + \text{H}_3\text{O}^+ + 2\text{e}^-$. How many grams of Pb will be used up to deliver 1 A for 100 h? ($\text{Pb} = 208$)
 (a) 776 g
 (b) 388 g
 (c) 194 g
 (d) 0.1 g
91. The copper anode of a cell containing silver nitrate solution weighs 60.0 g. After passing current for some time, it is found that 3.24 g of silver is deposited on the platinum cathode. What is the final weight of the anode? ($\text{Ag} = 108$, $\text{Cu} = 64$)
 (a) 0.96 g
 (b) 60 g
 (c) 59.04 g
 (d) 60.96 g
92. The quantity of electricity required for the reduction of 1 mole of Fe_2O_3 to Fe is
 (a) 1F
 (b) 0.33F
 (c) 3F
 (d) 6F
93. Three faradays of electricity is passed through molten Al_2O_3 , aqueous solutions of CuSO_4 and molten NaCl. The amounts of Al, Cu and Na deposited at the cathodes will be in the molar ratio of
 (a) 1:2:3
 (b) 3:2:1
 (c) 1:1.5:3
 (d) 6:3:2

94. Electrolysis of a solution of HSO_4^- ions produces $\text{S}_2\text{O}_8^{2-}$. Assuming 75% current efficiency, what current should be employed to achieve a production rate of 1 mole of $\text{S}_2\text{O}_8^{2-}$ per hour?
- (a) 71.5 A (b) 35.7 A
(c) 53.0 A (d) 143 A
95. The number of Faradays required to produce 1 g-atom of Mg from MgCl_2 is
- (a) 1 (b) 2
(c) 0.5 (d) 4
96. Passage of 96,500 coulomb of electricity liberates _____ L of O_2 at 273°C and 2 atm during electrolysis.
- (a) 5.6 (b) 16.8
(c) 22.4 (d) 11.2
97. An electrolytic cell contains a solution of Ag_2SO_4 and platinum electrodes. A current is passed until 1.6 g of O_2 has been liberated at anode. The amount of silver deposited at cathode would be
- (a) 108.0 g (b) 1.6 g
(c) 10.8 g (d) 21.6 g
98. When 12,000 coulombs of electricity is passed through the electrolyte, 3.0 g of a metal of atomic mass 96.5 g/mol is deposited. The electro-valency of the metal cation in the electrolyte is
- (a) +4 (b) +3
(c) +2 (d) -4
99. How many electrons flow when a current of 5 A is passed through a solution for 200 s?
- (a) 6.022×10^{23}
(b) 6.24×10^{21}
(c) 6.024×10^{21}
(d) 6.022×10^{20}
100. The current of 9.65 A flowing for 10 min deposits 3.0 g of a metal. The equivalent weight of the metal is
- (a) 10 (b) 30
(c) 50 (d) 96.5
101. If a current of 1.0 A is drawn from the Daniel cell for 96.5 min, the cathode will gain in weight by ($\text{Cu} = 63.5$, $\text{Zn} = 65.4$)
- (a) 1.905 g (b) 1.962 g
(c) 3.81 g (d) 3.924 g
102. The current required to produce oxygen at the rate of 2.8 ml per second during electrolysis of acidulated water is
- (a) 48.25 A/s
(b) 24.12 A/s
(c) 96.5 A/s
(d) 0.0048 A/s
103. Sodium amalgam is prepared by electrolysis of aqueous NaCl using 10 g mercury as cathode. How many Faraday of electricity is required to prepare 18.7% Na-amalgam, by weight, with a current efficiency of 50%?
- (a) 0.1 F (b) 0.2 F
(c) 0.05 F (d) 0.16 F
104. Element A (atomic mass = 112) and element B (atomic mass = 27) form chlorides. Solutions of these chlorides are electrolysed separately and it is found that when the same quantity of electricity is passed, 5.6 g of A was deposited while only 0.9 g of B was deposited. The valency of B is 3. The valency of A is
- (a) 1 (b) 2
(c) 3 (d) 4
105. The same current was passed successively through solution of zinc-ammonium sulphate and nickel-ammonium sulphate rendered alkaline with ammonia. The weights of zinc and nickel deposited in a certain time were found to be 22.89 g and 20.55 g, respectively. Given that the chemical equivalent weight of zinc is 32.7, what is the chemical equivalent weight of nickel?
- (a) 58.71 (b) 29.36
(c) 14.39 (d) 36.42
-

Conductance

- 106.** Which of the following solutions have highest resistance?
- 1N – NaCl
 - 0.05N – NaCl
 - 2N – NaCl
 - 0.1N – NaCl
- 107.** Variation of molar conductance of an electrolytic solution with temperature is that it
- increases with increase of temperature
 - decreases with increase of temperature
 - first increases then decreases
 - is not affected by temperature
- 108.** Which pure substance will not conduct electricity?
- Molten NaCl
 - Molten KOH
 - Liquefied HCl
 - Liquid Hg
- 109.** The correct order of molar conductance at infinite dilution of LiCl, NaCl and KCl is
- LiCl > NaCl > KCl
 - KCl > NaCl > LiCl
 - NaCl > KCl > LiCl
 - LiCl > KCl > NaCl
- 110.** The molar conductance of a strong electrolyte at infinite dilution
- tends to a finite value, which is above that at higher concentration
 - tends to a finite value, which is below that at higher concentration
 - tends to zero
 - tends to a finite value, which is equal to that at high concentration
- 111.** The best conductor of electricity is a 0.1 M solution of
- Boric acid
 - Sulphuric acid
 - Acetic acid
 - Propanoic acid
- 112.** The specific conductance of AgCl solution in water was determined to be $1.8 \times 10^{-6} \Omega^{-1} \text{ cm}^{-1}$ at 298 K. The molar conductances at infinite dilution, of Ag^+ and Cl^- are 67.9 and $82.1 \Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$, respectively. What is the solubility of AgCl in water?
- $1.2 \times 10^{-8} \text{ M}$
 - $1.44 \times 10^{-10} \text{ M}$
 - $1.2 \times 10^{-5} \text{ M}$
 - $1.44 \times 10^{-16} \text{ M}$
- 113.** Equivalence conductance at infinite dilution of NH_4Cl , NaOH and NaCl are 129.8, 217.4 and $108.9 \Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$, respectively. If the equivalent conductance of 0.01 N solution of NH_4OH is $9.532 \Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$, then the degree of dissociation of NH_4OH at this temperature is
- 0.04%
 - 2.1%
 - 4.0%
 - 44.7%
- 114.** The resistance of 1 M – CH_3COOH solution is 250Ω , when measured in a cell of cell constant 125 m^{-1} . The molar conductivity, in $\Omega^{-1} \text{ m}^2 \text{ mol}^{-1}$ is
- 5.0×10^{-4}
 - 500
 - 2×10^{-3}
 - 200
- 115.** How does the electrical conductivity of 20 ml of 0.2 M – MgSO_4 change when 0.5 M – Ba(OH)_2 solution is gradually added in it, to excess?
- decreases continuously
 - increases continuously
 - increases and then decreases
 - decreases and then increases
- 116.** The equivalent conductivity (in $\Omega^{-1} \text{ cm}^2 \text{ eq}^{-1}$) of 1.0 M – H_2SO_4 solution of specific conductance $2.6 \times 10^{-1} \text{ cm}^{-1}$, is
- 1.3×10^2
 - 6.5×10^1
 - 1.3×10^{-1}
 - 2.6×10^2

- 117.** The molar conductance of a 0.01 M solution of acetic acid was found to be $16.30 \Omega^{-1} \text{ cm}^{-1} \text{ mol}^{-1}$ at 25°C . The ionic conductances of hydrogen and acetate ions at infinite dilution are 349.8 and $40.9 \Omega^{-1} \text{ cm}^{-1} \text{ mol}^{-1}$, respectively, at the same temperature. What percentage of acetic acid is dissociated at this concentration?
- (a) 0.04172% (b) 4.172%
(c) 41.72% (d) 0.4172%
- 118.** The distance between two electrodes of a cell is 2.5 cm and area of each electrode is 5 cm^2 . The cell constant is
- (a) 0.5 m^{-1} (b) 12.5 cm^3
(c) 2.0 cm (d) 50 m^{-1}
- 119.** The molar conductivity of NH_4Cl , OH^- and Cl^- at infinite dilution is 150, 200 and $75 \Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$, respectively. If the molar conductivity of a 0.01 M- NH_4OH solution is $22 \Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$, then its degree of dissociation is
- (a) 0.146 (b) 0.063
(c) 0.080 (d) 0.293
- 120.** Calculate the ionic product of water at 25°C from the following data:
- Conductivity of water = $5.5 \times 10^{-6} \text{ mho m}^{-1}$
 $\lambda_{\text{H}^+}^\circ = 0.035 \text{ mho m}^2 \text{ mol}^{-1}$
 $\lambda_{\text{OH}^-}^\circ = 0.020 \text{ mho m}^2 \text{ mol}^{-1}$
- (a) $2 \times 10^{-14} \text{ M}^2$
(b) $1 \times 10^{-7} \text{ M}^2$
(c) $1 \times 10^{-8} \text{ M}^2$
(d) $1 \times 10^{-14} \text{ M}^2$
- 121.** Calculate K_a of acetic acid if its 0.05 M solution has molar conductivity of $7.814 \times 10^{-4} \Omega^{-1} \text{ m}^2 \text{ mol}^{-1}$ at 25°C . Given: Λ_m° for $\text{CH}_3\text{COOH} = 3.907 \times 10^{-2} \Omega^{-1} \text{ m}^2 \text{ mol}^{-1}$.
- (a) 2×10^{-5} (b) 1.8×10^{-5}
(c) 4×10^{-4} (d) 0.02
- 122.** Equal volumes of 0.015 M – CH_3COOH and 0.015 M – NaOH solutions are mixed together. What would be the molar conductivity of mixture if conductivity of CH_3COONa is $6.3 \times 10^{-4} \text{ S cm}^{-1}$?
- (a) $0.84 \text{ S cm}^2 \text{ mol}^{-1}$
(b) $8.4 \text{ S cm}^2 \text{ mol}^{-1}$
(c) $84 \text{ S cm}^2 \text{ mol}^{-1}$
(d) $42 \text{ S cm}^2 \text{ mol}^{-1}$
- 123.** Calculate Λ_m^∞ (in $\Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$) for SrCl_2 at 25°C , from the following data:
- | Conc. | 0.25 M | 1.0 M |
|---|--------|-------|
| Λ_m (in $\Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$) | 260 | 250 |
- (a) 270 (b) 265
(c) 240 (d) 275
- 124.** The resistance of a solution A is 50Ω and that of solution B is 100Ω , both solutions being taken in the same conductivity cell. If equal volumes of solution A and B are mixed, what will be the resistance of the mixture using the same cell? Assume that there is no increase in the degree of dissociation of A and B on mixing.
- (a) 150Ω
(b) 75Ω
(c) 33.33Ω
(d) 66.67Ω
- 125.** In a conductivity cell, the two platinum electrodes, each of area 10 cm^2 are fixed 1.5 cm apart. The cell contained 0.05 N solution of a salt. If the two electrodes are just half dipped into the solution which has a resistance of 50Ω , the equivalent conductance of the salt solution, in $\Omega^{-1} \text{ cm}^2 \text{ eq}^{-1}$, is
- (a) 120
(b) 60
(c) 240
(d) 3000

Answer Keys – Exercise I**Electrode Potential**

1. (c) 2. (d) 3. (d) 4. (a) 5. (a) 6. (a) 7. (b) 8. (a) 9. (a) 10. (a)
11. (d) 12. (b) 13. (a) 14. (c) 15. (c) 16. (b) 17. (b) 18. (a) 19. (c) 20. (a)
21. (c) 22. (c) 23. (c) 24. (a) 25. (d) 26. (d) 27. (c) 28. (c) 29. (a) 30. (a)

Galvanic Cell

31. (a) 32. (d) 33. (b) 34. (b) 35. (c) 36. (c) 37. (c) 38. (c) 39. (d) 40. (d)
41. (d) 42. (d) 43. (c) 44. (b) 45. (b) 46. (d) 47. (d) 48. (a) 49. (d) 50. (c)
51. (b) 52. (b) 53. (a) 54. (a) 55. (a) 56. (c) 57. (a) 58. (c) 59. (c) 60. (b)
61. (a) 62. (b) 63. (c) 64. (d) 65. (a)

Electrolysis

66. (b) 67. (a) 68. (a) 69. (b) 70. (b) 71. (d) 72. (b) 73. (c) 74. (b) 75. (d)
76. (a) 77. (b) 78. (d) 79. (b) 80. (a) 81. (b) 82. (a) 83. (c) 84. (d) 85. (a)
86. (a) 87. (d) 88. (a) 89. (c) 90. (b) 91. (c) 92. (d) 93. (c) 94. (a) 95. (b)
96. (a) 97. (c) 98. (a) 99. (b) 100. (c) 101. (a) 102. (a) 103. (b) 104. (b) 105. (b)

Conductance

106. (b) 107. (a) 108. (c) 109. (b) 110. (a) 111. (b) 112. (c) 113. (c) 114. (a) 115. (d)
116. (a) 117. (b) 118. (d) 119. (c) 120. (b) 121. (a) 122. (c) 123. (a) 124. (d) 125. (a)
-

EXERCISE II (JEE ADVANCED)

Section A (Only one Correct)

1. A student made the following observations in the laboratory:

- (I) Clean copper metal did not react with 1 M – $\text{Pb}(\text{NO}_3)_2$ solution.
 (II) Clean lead metal dissolves in 1 M – AgNO_3 solution and crystals of Ag metal appeared.
 (III) Clean silver metal did not react with 1 M – $\text{Cu}(\text{NO}_3)_2$ solution.

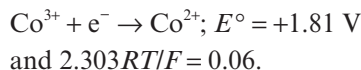
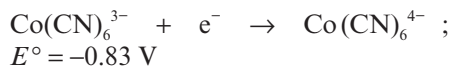
The order of decreasing reducing character of the three metals is

- (a) $\text{Cu} > \text{Pb} > \text{Ag}$
 (b) $\text{Cu} > \text{Ag} > \text{Pb}$
 (c) $\text{Pb} > \text{Cu} > \text{Ag}$
 (d) $\text{Pb} > \text{Ag} > \text{Cu}$
2. We have an oxidation–reduction system: $[\text{Fe}(\text{CN})_6]^{3-} + \text{e}^- \rightleftharpoons [\text{Fe}(\text{CN})_6]^{4-}$; $E^\circ = +0.36 \text{ V}$. The ratio of concentrations of oxidized and reduced form at which the potential of the system becomes 0.24 V, is [Given: $2.303RT/F = 0.06$]
- (a) 2:1 (b) 1:2
 (c) 1:20 (d) 1:100
3. The standard reduction potential for the process: $[\text{Co}(\text{H}_2\text{O})_6]^{3+} + \text{e}^- \rightarrow [\text{Co}(\text{H}_2\text{O})_6]^{2+}$ is 1.8 V. The standard reduction potential for the process: $[\text{Co}(\text{NH}_3)_6]^{3+} + \text{e}^- \rightarrow [\text{Co}(\text{NH}_3)_6]^{2+}$ is 0.1 V. Which of the complex ion, $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ or $[\text{Co}(\text{NH}_3)_6]^{2+}$ can be oxidized to the corresponding cobalt (III) complex, by oxygen, in basic medium, under standard condition? [Given: $E^\circ_{\text{O}_2/\text{OH}^-} = 0.4 \text{ V}$]
- (a) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$
 (b) $[\text{Co}(\text{NH}_3)_6]^{2+}$
 (c) both
 (d) none of these

4. The standard reduction potential for the reactions: $\text{Ag}^+ + \text{e}^- \rightarrow \text{Ag}$ and $\text{Ag}(\text{NH}_3)_2^+ + \text{e}^- \rightarrow \text{Ag} + 2\text{NH}_3$ are +0.79 V and +0.37 V, respectively. From these values and the Nernst equation, what should be K_f for the $\text{Ag}(\text{NH}_3)_2^+$ ion? [Given: $2.303RT/F = 0.06$]

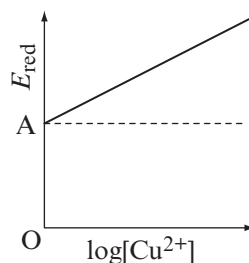
- (a) 1.0×10^{-7}
 (b) 1.0×10^7
 (c) 2.15×10^{19}
 (d) 4.64×10^{-20}

5. The overall formation constant for the reaction of 6 mole of CN^- with cobalt (II) is 1×10^{19} . What is the formation constant for the reaction of 6 moles of CN^- with cobalt (III)? Given that



- (a) 1.0×10^{63} (b) 1.0×10^{25}
 (c) 1.0×10^{-25} (d) 1.0×10^{-63}

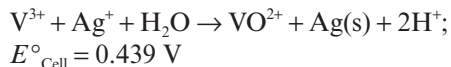
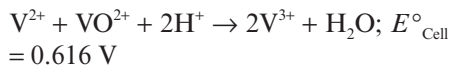
6. For the process: $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$; $\log[\text{Cu}^{2+}]$ vs. E_{red} graph is shown in the figure, where $\text{OA} = 0.34 \text{ V}$. The electrode potential of the half-cell of $\text{Cu}|\text{Cu}^{2+}$ (0.1 M) will be [$2.303RT/F = 0.06$]



- (a) -0.31 V (b) +0.31 V
 (c) -0.37 V (d) +0.37 V

7. Tell which of the following statements accurately describes the effect of adding CN^- to the cathode of a cell with a cell reaction: $\text{Cd} + 2\text{Ag}^+ \rightarrow 2\text{Ag} + \text{Cd}^{2+}$, $E^\circ = 1.2 \text{ V}$
- E° increases because $\text{Cd}(\text{CN})_4^{2-}$ forms
 - E° decreases because $\text{Cd}(\text{CN})_4^{2-}$ forms
 - E° increases because $\text{Ag}(\text{CN})_2^-$ forms
 - E° decreases because $\text{Ag}(\text{CN})_2^-$ forms
8. For an electrochemical cell $\text{Zn} \mid \text{Zn}^{2+} (C_1 \text{ M}) \parallel \text{Cu}^{2+} (C_2 \text{ M}) \mid \text{Cu}$, the decrease in free energy at a given temperature is a function of
- $\ln C_1$
 - $\ln C_2$
 - $\ln C_2 \cdot C_1$
 - $\ln C_1/C_2$
9. Select the correct option if it is known that $K_{\text{sp}}(\text{AgCl}) > K_{\text{sp}}(\text{AgBr}) > K_{\text{sp}}(\text{AgI})$
- $E^\circ_{\text{I}^- \mid \text{AgI} \mid \text{Ag}} > E^\circ_{\text{Br}^- \mid \text{AgBr} \mid \text{Ag}} > E^\circ_{\text{Cl}^- \mid \text{AgCl} \mid \text{Ag}}$
 - $E^\circ_{\text{I}^- \mid \text{AgI} \mid \text{Ag}} < E^\circ_{\text{Br}^- \mid \text{AgBr} \mid \text{Ag}} < E^\circ_{\text{Cl}^- \mid \text{AgCl} \mid \text{Ag}}$
 - $E^\circ_{\text{I}^- \mid \text{AgI} \mid \text{Ag}} < E^\circ_{\text{Cl}^- \mid \text{AgCl} \mid \text{Ag}} < E^\circ_{\text{Br}^- \mid \text{AgBr} \mid \text{Ag}}$
 - $E^\circ_{\text{I}^- \mid \text{AgI} \mid \text{Ag}} = E^\circ_{\text{Br}^- \mid \text{AgBr} \mid \text{Ag}} = E^\circ_{\text{Cl}^- \mid \text{AgCl} \mid \text{Ag}}$
10. The EMF of a galvanic cell composed of two hydrogen electrodes is 272 mV. What is the pH of the solution in which the anode is immersed if the cathode is in contact with a solution of pH = 3?
- 3
 - 6.7
 - 7.6
 - 1.6
11. At what pH does the potential (EMF) for the disproportionation of chlorine change from a negative value to a positive value, assuming 1.0 bar pressure and 1.0 M concentration for all species except hydrogen ion? Given:
- $$\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-; E^\circ = 1.36 \text{ V}$$
- $$2\text{OCl}^- + 4\text{H}^+ + 2\text{e}^- \rightarrow \text{Cl}_2 + 2\text{H}_2\text{O}; E^\circ = 1.63 \text{ V}$$
- [$2.303RT/F = 0.06$]
- 4.5
 - 1.5
 - 2.25
 - 9.0
12. The EMF of the cell: $\text{Hg}(\text{l}) \mid \text{Hg}_2\text{Cl}_2(\text{s}), \text{KCl sol. (1.0N)} \mid \text{Quinohydrone} \mid \text{Pt}$, is 0.210 V at 298 K. What is the pH of the quinohydrone solution, the potential of the normal calomel electrode is 0.279 V and E° for the quinohydrone electrode is 0.699 V, both at the same temperature. [$2.303RT/F = 0.06$]
- 3.5
 - 7.0
 - 1.85
 - 3.5
13. The potential (EMF) of a cell consisting of an anode of silver in 0.10 M AgNO_3 solution and a cathode of Pt immersed in a solution of 1.6 M $\text{Cr}_2\text{O}_7^{2-}$, 0.4 M Cr^{3+} and 0.1 M H^+ is ($E^\circ_{\text{Ag}^+ \mid \text{Ag}} = 0.80 \text{ V}$ and $E^\circ_{\text{Cr}_2\text{O}_7^{2-} \mid \text{Cr}^{3+}} = 1.33 \text{ V}$)
- 0.46 V
 - 0.60 V
 - 0.53 V
 - 0.17 V
14. The EMF of the cell: $\text{Zn-Hg} (C_1 \text{ M}) \mid \text{Zn}^{2+}(\text{aq}) \mid \text{Zn-Hg} (C_2 \text{ M})$ at 25°C , if the concentrations of the zinc amalgams are: $C_1 = 10 \text{ g per } 100 \text{ g of mercury}$, $C_2 = 1 \text{ g per } 100 \text{ g mercury}$, is
- 0.059 V
 - 0.0295 V
 - 0.59 V
 - 0.295 V
15. What is the equilibrium constant of the reaction: $2\text{Fe}^{3+} + \text{Au}^+ \rightarrow 2\text{Fe}^{2+} + \text{Au}^{3+}$? Given $E^\circ_{\text{Au}^+ \mid \text{Au}} = 1.68 \text{ V}$, $E^\circ_{\text{Au}^{3+} \mid \text{Au}} = 1.50 \text{ V}$, $E^\circ_{\text{Fe}^{3+} \mid \text{Fe}^{2+}} = 0.75 \text{ V}$ and $2.303RT/F = 0.06$.
- 1×10^{22}
 - 1×10^{-22}
 - 1×10^{-11}
 - 1×10^{-72}
16. What is the EMF of the cell: $\text{Pt}, \text{H}_2 (1 \text{ atm}) \mid \text{CH}_3\text{COOH} (0.1 \text{ M}) \parallel (0.01 \text{ M}) \text{NH}_4\text{OH} \mid \text{H}_2 (1 \text{ atm}), \text{Pt}$? Given: K_a for $\text{CH}_3\text{COOH} = 1.8 \times 10^{-5}$, K_b for $\text{NH}_4\text{OH} = 1.8 \times 10^{-5}$, $2.303RT/F = 0.06$, $\log 1.8 = 0.25$
- 0.465 V
 - 0.465 V
 - 0.2325 V
 - 0.93 V

17. Two electrochemical cells are assembled in which the following reactions occur:



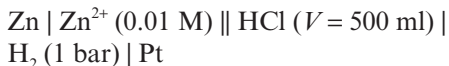
If $E^\circ_{\text{Ag}^+|\text{Ag}} = 0.799 \text{ V}$, what is $E^\circ_{\text{V}^{3+}|\text{V}^{2+}}$?

- (a) -0.256 V (b) $+0.256 \text{ V}$
(c) $+1.854 \text{ V}$ (d) -1.854 V
18. The EMF of cell: $\text{H}_2(\text{g}) | \text{Buffer} || \text{Normal calomel electrode}$, is 0.70 V at 25°C , when barometric pressure is 760 mm . What is the pH of the buffer solution? $E^\circ_{\text{Calomel}} = 0.28 \text{ V}$. [$2.303RT/F = 0.06$]
- (a) 3.5
(b) 7.0
(c) tending to zero
(d) tending to 14.0
19. What is the solubility product of a saturated solution of Ag_2CrO_4 in water at 298 K if the EMF of the cell: $\text{Ag} | \text{Ag}^+ (\text{satd. } \text{Ag}_2\text{CrO}_4) || \text{Ag}^+ (0.1 \text{ M}) | \text{Ag}$ is 0.162 V at 298 K ? [$2.303RT/F = 0.06$, $\log 2 = 0.3$]
- (a) 2.0×10^{-4} (b) 3.2×10^{-11}
(c) 8.0×10^{-12} (d) 4.0×10^{-12}
20. A $\text{Tl}^+|\text{Tl}$ couple was prepared by saturating 0.10 M KBr with TlBr and allowing Tl^+ ions form the insoluble bromide to equilibrate. This couple was observed to have a potential -0.444 V with respect to $\text{Pb}^{2+} | \text{Pb}$ couple in which Pb^{2+} was 0.10 M . What is the K_{sp} of TlBr . [Given: $E^\circ_{\text{Pb}^{2+}|\text{Pb}} = -0.126 \text{ V}$, $E^\circ_{\text{Tl}^+|\text{Tl}} = -0.336 \text{ V}$, $\log 2.5 = 0.4$, $2.303RT/F = 0.06$]
- (a) 4.0×10^{-6} (b) 2.5×10^{-4}
(c) 4.0×10^{-5} (d) 6.3×10^{-3}
21. The EMFs of the cell obtained by combining separately Zn and Cu electrodes of a Daniel cell with normal calomel electrodes are 1.083 V and

-0.018 V , respectively, at 25°C . If the potential of normal calomel electrode is -0.28 V , the EMF of the Daniel cell is

- (a) 1.065 V (b) 1.101 V
(c) 0.803 V (d) 0.262 V
22. A hydrogen electrode placed in a buffer solution of CH_3COONa and CH_3COOH in the ratio's $x:y$ and $y:x$ has electrode potential values E_1 and E_2 volts, respectively, at 25°C . The $\text{p}K_a$ value of acetic acid is
- (a) $(E_1 + E_2)/0.118$
(b) $(E_2 - E_1)/0.118$
(c) $-(E_1 + E_2)/0.118$
(d) $(E_1 - E_2)/0.118$
23. The standard reduction potential at 25°C of the reaction $2\text{H}_2\text{O} + 2\text{e}^- \rightleftharpoons \text{H}_2 + 2\text{OH}^-$ is -0.84 V . Calculate equilibrium constant for the reaction: $2\text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ at 25°C . ($2.303RT/F = 0.06$)
- (a) 10^{-14} (b) 10^{14}
(c) 10^7 (d) 10^{-7}
24. The voltage of the cell given below is -0.61 V .
- $$\text{Pt} | \text{H}_2 (1 \text{ bar}) | \text{NaHSO}_3 (0.4 \text{ M}), \text{Na}_2\text{SO}_3 (6.4 \times 10^{-2} \text{ M}) || \text{Zn}^{2+} (0.4 \text{ M}) | \text{Zn}$$
- If $E^\circ_{\text{Zn}^{2+}|\text{Zn}} = -0.76 \text{ V}$, Calculate K_{a2} of H_2SO_4 . ($2.303RT/F = 0.06$)
- (a) 3.2×10^{-4} (b) 3.2×10^{-2}
(c) 3.2×10^{-3} (d) 6.4×10^{-7}
25. Estimate the cell potential of a Daniel cell having 1.0 M Zn^{2+} and originally having 1.0 M Cu^{2+} after sufficient ammonia has been added to the cathode compartment to make the NH_3 concentration 2.0 M . Given: $E^\circ_{\text{Zn}^{2+}|\text{Zn}} = -0.76 \text{ V}$, $E^\circ_{\text{Cu}^{2+}|\text{Cu}} = +0.34 \text{ V}$, K_f for $\text{Cu}(\text{NH}_3)_4^{2+} = 1 \times 10^{12}$.
- (a) 1.10 V
(b) 0.704 V
(c) 0.396 V
(d) 1.496 V

26. The minimum mass of NaOH required to be added in RHS to consume all the H^+ present in RHS of the cell of EMF, +0.70 V at 25°C, before its use.

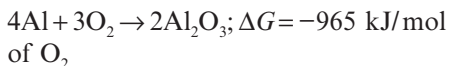


Given: $E^\circ_{Zn^{2+}|Zn} = -0.76 \text{ V}$; $2.303RT/F = 0.06$

- (a) 2.0 g (b) 4.0 g
(c) 0.2 g (d) 0.4 g
27. The EMF of the cell: Ag, AgCl in 0.1 M – KCl || satd. NH_4NO_3 || 0.1 M – $AgNO_3$, Ag is 0.42 V at 25°C. 0.1 M – KCl is 50% dissociated and 0.1 M – $AgNO_3$ is 40% dissociated. The solubility product of AgCl is ($2.303RT/F = 0.06$)

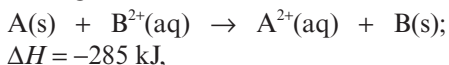
- (a) 1.0×10^{-10}
(b) 2.0×10^{-9}
(c) 1.0×10^{-9}
(d) 2.0×10^{-10}

28. On the basis of information available from the reaction



The minimum EMF required to carry out electrolysis of Al_2O_3 is

- (a) 0.833 V (b) 2.5 V
(c) 5.0 V (d) 1.67 V
29. The theoretical efficiency of a hypothetical cell is about 84% which involves the following reaction



then, the standard EMF of the cell is

- (a) 1.10 V (b) 1.24 V
(c) 2.48 V (d) 2.20 V
30. To a beaker containing 0.1 M – HCl, a little pure solid AgCl is added. Both a silver and a hydrogen electrode ($P_{H_2} = 1.0 \text{ bar}$) are then placed in the solution. What is the approximate value of EMF

measured between the electrodes at 25°C? (Given: $E^\circ_{Ag^+|Ag} = 0.80 \text{ V}$ and $K_{sp} (AgCl) = 10^{-10}$)

- (a) 0.92 V (b) 0.32 V
(c) 1.28 V (d) 0.56 V

31. Although aluminium is above hydrogen in the electrochemical series, it is stable in air and water. Why?

- (a) Aluminium is non-reactive metal.
(b) A layer of oxide is formed at metal surface, which prevent further reaction.
(c) The reaction is thermodynamically unfavourable, i.e., ΔG is not negative.
(d) Kinetically, the reaction has very low activation energy and hence do not occur.

32. When silver chloride is dissolved in a large excess of ammonia, practically all silver ion can be assumed to exist in form of a single ionic species $[Ag_x(NH_3)_y]^{x+}$. Compute the values of x and y using the two following cells:

Cell I: Ag | $4.0 \times 10^{-4} \text{ M}$ – AgCl, 1 M – NH_3 || $4 \times 10^{-2} \text{ M}$ – AgCl, 1 M – NH_3 | Ag; $E_{cell} = 0.118 \text{ V}$ at 298 K.

Cell II: Ag | $3 \times 10^{-3} \text{ M}$ – AgCl, 1 M – NH_3 || $3 \times 10^{-3} \text{ M}$ – AgCl, 0.1 M – NH_3 | Ag; $E_{cell} = 0.118 \text{ V}$ at 298 K.

- (a) $x = 1, y = 2$
(b) $x = 1, y = 4$
(c) $x = 2, y = 4$
(d) $x = 1, y = 6$

33. When metallic copper is shaken with a solution of a copper salt, the reaction $Cu + Cu^{2+} \rightleftharpoons 2Cu^+$ proceeds. When equilibrium is established at 298 K, $[Cu^{2+}]/[Cu^+]^2 = 1.667 \times 10^6 \text{ M}^{-1}$. If the standard potential of the $Cu^{2+} | Cu$ half-cell is +0.3376 V, what is the standard potential of $Cu^+ | Cu$ half-cell? (Given: $2.303RT/F = 0.06$, $\log 2 = 0.3$, $\log 3 = 0.48$)

- (a) -0.3732 V (b) 0.6752 V
(c) 0.5242 V (d) 0.151 V

34. Zinc granules are added in excess to a 500 ml of 1.0 M nickel nitrate solution at 25°C until the equilibrium is reached. If the standard reduction potential of $\text{Zn}^{2+}|\text{Zn}$ and $\text{Ni}^{2+}|\text{Ni}$ are -0.75 V and -0.24 V , respectively, the concentration of Ni^{2+} in solution at equilibrium is ($2.303RT/F = 0.06$)

(a) $1.0 \times 10^{-17}\text{ M}$
 (b) $1.0 \times 10^{17}\text{ M}$
 (c) $5 \times 10^{-17}\text{ M}$
 (d) $2 \times 10^{-17}\text{ M}$

35. The EMF for the cell: $\text{Ag(s)} | \text{AgCl(s)} | \text{KCl(0.2 M)} || \text{KBr(0.001 M)} | \text{AgBr(s)} | \text{Ag(s)}$ at 25°C is ($K_{\text{sp}}(\text{AgCl}) = 2.0 \times 10^{-10}$; $K_{\text{sp}}(\text{AgBr}) = 4.0 \times 10^{-13}$, $2.303RT/F = 0.06$, $\log 2 = 0.3$)

(a) 0.024 V
 (b) -0.024 V
 (c) -0.24 V
 (d) -0.012 V

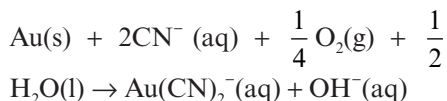
36. Two students use same stock solution of ZnSO_4 but different solutions of CuSO_4 . The EMF of one cell is 0.03 V higher than the other. The concentration of CuSO_4 in the cell with higher EMF value is 0.5 M. The concentration of CuSO_4 in the other cell is ($2.303RT/F = 0.06$)

(a) 0.05 M (b) 5.0 M
 (c) 0.5 M (d) 0.005 M

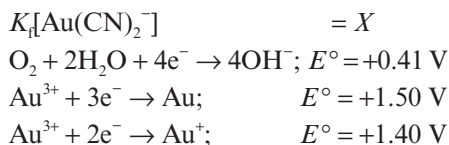
37. Two weak acid solutions HA_1 and HA_2 each with the same concentration and having $\text{p}K_{\text{a}}$ values 3 and 5 are placed in contact with hydrogen electrodes (1 atm, 25°C) and are interconnected through a salt bridge. EMF of the cell is

(a) 0.0295 V (b) 0.118 V
 (c) 0.0885 V (d) 0.059 V

38. Consider the reaction of extraction of gold from its ore:



Use the following data to calculate ΔG° for the above reaction.

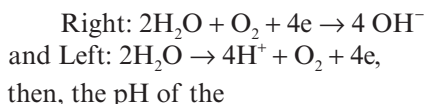


(a) $-RT \ln X + 1.29\text{ F}$
 (b) $-RT \ln X - 1.29\text{ F}$
 (c) $+RT \ln X + 2.11\text{ F}$
 (d) $-RT \ln X - 2.11\text{ F}$

39. F_2 gas cannot be obtained by the electrolysis of any aqueous fluoride salt because

(a) F_2 is the strongest oxidizing agent
 (b) F_2 easily combines with water
 (c) F_2 readily combines with the electrodes
 (d) F^- can never be oxidized

40. A volume of 100 ml of a buffer of 1 M $-\text{NH}_3$ and 1 M $-\text{NH}_4^+$ is placed in two half-cells connected by a salt bridge. A current of 1.5 A is passed through the cell for 20 min. If electrolysis of water takes place only and the electrode reactions are:



(a) right electrode will increase
 (b) left electrode will increase
 (c) both electrode will increase
 (d) both electrode will decrease

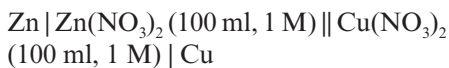
41. When molten ICl_3 is electrolysed using platinum electrodes

(a) I_2 is evolved at cathode and Cl_2 at anode
 (b) Cl_2 is evolved at cathode and I_2 at anode
 (c) I_2 is evolved at cathode and both I_2 and Cl_2 at anode
 (d) electrolysis does not take place

42. When electric current is passed through a cell having an electrolyte, positive ions move towards the cathode and negative ions towards the anode. If the cathode is pulled out of the solution, then the
- positive and negative ions will move towards the anode
 - positive ions will start moving towards the anode, the negative ions will stop moving
 - negative ions will continue to move towards the anode, the positive ions will start moving randomly
 - positive and negative ions will start moving randomly
43. When 10^{-6} M – HCl is electrolysed
- O_2 is produced at the anode
 - H_2 is produced at the anode
 - Cl_2 is produced at the anode
 - Cl_2 and O_2 are produced at the anode
44. A dilute aqueous solution of Na_2SO_4 is electrolysed using platinum electrodes. The products at the anode and cathode are, respectively,
- O_2, H_2
 - $S_2O_8^{2-}, Na$
 - O_2, Na
 - $S_2O_8^{2-}, H_2$
45. In the electrolysis of a fused salt, the mass of substance deposited on an electrode will not depend on
- temperature of bath
 - current intensity
 - time of electrolysis
 - electrochemical equivalent of ions
46. When an electric current is passed through an aqueous solution of the following, the concentrations of cation as well as anion will not change for (assume constant volume of the electrolytic solution)
- CsCl
 - KNO_3
 - $AgNO_3$
 - HCl
47. Which of the following substances: Na, Hg, S, Pt and graphite can be used as electrodes in electrolytic cells having aqueous solution?
- Na, Pt and graphite
 - Na and Hg
 - Hg, Pt and graphite
 - Na and S
48. An ammeter and a copper voltameter are connected in series in an electric circuit through which a constant direct current flows. The ammeter shows 0.5 A. If 0.635 g of Cu is deposited in 0.965 h, what is the percentage error of ammeter? (Cu = 63.5)
- 5 %
 - 10 %
 - 9 %
 - 90 %
49. The same current is passed through acidulated water and stannous chloride solution. What volume of dry detonating gas at $0^\circ C$ and 1 atm is evolved from water, when 1.20 g of tin is deposited from the other solution? (Sn = 120)
- 112 ml
 - 336 ml
 - 224 ml
 - 672 ml
50. The preparation of LiOH by the electrolysis of a 35% solution of LiCl using a platinum anode led to a current efficiency of 80%. What weight of LiOH was formed by the passage of 2.5 A for 4825 s?
- 1.92 g
 - 2.40 g
 - 0.96 g
 - 0.672 g
51. Assuming that copper contains only iron, silver and gold as impurities. After passage of 12.4 A for 4825 s, the mass of anode decreased by 20.00 g and the cathode increased by 19.05 g. The percentages of iron and copper in the original sample are, respectively, (Cu = 63.5, Fe = 56)
- 4.75%, 95.25%
 - 2.8%, 95.25%
 - 95.25%, 4.75%
 - 95.25%, 1.91%

52. After electrolysis of an aqueous sodium chloride solution, it was found that the solution is being neutralized by 60 ml N – HCl solution. During the same period of electrolysis, 3.18 g of copper was deposited in a copper voltameter in series. What is the percentage of the theoretical yield of sodium hydroxide obtained? (Cu = 63.6)
- (a) 40% (b) 60%
(c) 30% (d) 80%
53. A lead storage battery has initially 200 g of lead and 200 g of PbO_2 plus excess H_2SO_4 . Theoretically, how long could this cell deliver a current of 10.0 A, without recharging, if it were possible to operate it so that the reaction goes to completion? (Pb = 208)
- (a) 16083.33 s (b) 8041.67 s
(c) 44.68 s (d) 18557.7 s
54. A volume of 100 ml of 0.6 N- CuSO_4 solution is electrolysed between two platinum electrodes till the concentration in the residual liquid is 0.1 N, when a steady current of 5.0 A is used. How long should the current be passed to get the above change?
- (a) 965 s (b) 96500 s
(c) 1930 s (d) 482.5 s
55. A constant current flowed for 2 h through a potassium iodide solution, oxidizing the iodide ion to iodine. At the end of the experiment, the iodine was titrated with 72 ml of 1.0 M- $\text{Na}_2\text{S}_2\text{O}_3$ solution. What was the average rate of current flow, in ampere?
- (a) 0.965 A (b) 1.93 A
(c) 0.483 A (d) 0.0965 A
56. Anthracene, $\text{C}_{14}\text{H}_{10}$, can be oxidized to anthraquinone, $\text{C}_{14}\text{H}_8\text{O}_2$. What weight of anthraquinone can be produced by the passage of a current of 1 A of 40 min if the current efficiency is 96.5%?
- (a) 0.862 g (b) 0.832 g
(c) 0.624 g (d) 0.738 g
57. Most of the copper used to make wire has been electrically refined by depositing it from copper salts solution (divalent) on to a cathode. What is the cost of electrical energy required per kg of copper if the cost of electricity is Rs 4.00 per kWh and the cell operates at 0.33 V? The electrochemical equivalent of copper is 0.00033 g/coulomb.
- (a) Rs 11.11 (b) Rs 5.55
(c) Rs 2.22 (d) Rs 1.11
58. Electrolysis of an acetate solution produces ethane according to the reaction:
- $$2\text{CH}_3\text{COO}^- \rightarrow \text{C}_2\text{H}_6(\text{g}) + 2\text{CO}_2(\text{g}) + 2\text{e}^-$$
- What total volume of ethane and CO_2 would be produced at 0°C and 1 atm, if a current of 0.5 A is passed through the solution for 482.5 min? Assume current efficiency 80%.
- (a) 1.344 L (b) 2.688 L
(c) 4.032 L (d) 1.792 L
59. To perform an analysis of a mixture of metal ions by electro-deposition, the second metal to be deposited must not be plating out until the concentration ratio of the second to the first is about 10^6 . What must be the minimum difference in standard potential of the two metals which form dipositive ions in order for such an analysis to be feasible?
- (a) 0.177 V (b) 0.354 V
(c) 0.708 V (d) 0.088 V
60. During the electrolysis of 0.1 M – CuSO_4 solution using copper electrodes, a depletion of Cu^{2+} occurs near the cathode with a corresponding excess near the anode, owing to inefficient stirring of the solution. If the local concentration of Cu^{2+} near the anode and cathode are, respectively, 0.12 M and 0.08 M, the back EMF developed at 298 K is ($\log 1.5 = 0.18$, $2.303RT/F = 0.06$)
- (a) 0.33 V (b) 5.4 mV
(c) 2.7 mV (d) 10.8 mV

61. The following galvanic cell:



was operated as an electrolytic cell as Cu as the anode and Zn as the cathode. A current of 0.4825 A was passed for 10 h and then the cell was allowed to function as galvanic cell. What would be the final EMF of the cell at 25°C? Assume that the only electrode reactions occurring were those involving $\text{Cu} | \text{Cu}^{2+}$ and $\text{Zn} | \text{Zn}^{2+}$. Given: $E^\circ_{\text{Cu}^{2+}|\text{Cu}} = +0.34 \text{ V}$, $E^\circ_{\text{Zn}^{2+}|\text{Zn}} = -0.76 \text{ V}$, $2.303RT/F = 0.06$, $\log 1.9 = 0.28$

- (a) 1.10 V (b) 1.0616 V
(c) 1.1384 V (d) 1.1768 V

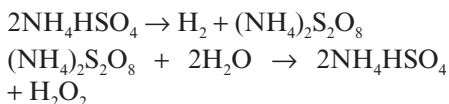
62. At the Nangal fertilizer plant in Punjab, hydrogen is produced by the electrolysis of water. The hydrogen is used for the production of ammonia and nitric acid (by the oxidation of ammonia). If the average production of ammonium nitrate is 1152 kg/day, the daily consumption of electricity (in A /day) is

- (a) 96,500 (b) 48,250
(c) 32,166.67 (d) 24,125

63. Lactic acid, $\text{HC}_3\text{H}_5\text{O}_3$, produced in 1 g sample of muscle tissue was titrated using phenolphthalein as indicator against OH^- ions which were obtained by the electrolysis of water. As soon as OH^- ions are produced, they react with lactic acid and at complete neutralization, immediately a pink colour is noticed. If electrolysis was made for 1158 s using 50.0 mA current to reach the end point, what was the percentage of lactic acid in muscle tissue?

- (a) 5.4% (b) 2.7%
(c) 10.8% (d) 0.054%

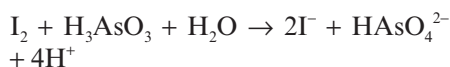
64. H_2O_2 can be prepared by successive reactions:



The first reaction is an electrolytic reaction and second is steam distillation. What amount of current would have to be used in first reaction to produce enough intermediate to yield 102 g pure H_2O_2 per hour. Assume current efficiency 50%.

- (a) 643.33 A
(b) 321.67 A
(c) 160.83 A
(d) 1286.67 A

65. In an analytical determination of arsenic, a solution containing arsenious acid, H_3AsO_3 , KI and a small amount of starch is electrolysed. The electrolysis produces free I_2 from I^- ions and the I_2 immediately oxidizes the arsenious acid to hydrogen arsenate ion, HAsO_4^{2-} .



When the oxidation of arsenic is complete, the free iodine combines with the starch to give a deep blue colour. If during a particular run, it takes 96.5 s for a current of 10.0 mA to give an end point (indicated by the blue colour), how many g of arsenic are present in the solution. (As = 75)

- (a) $6.3 \times 10^{-4} \text{ g}$
(b) $1.26 \times 10^{-4} \text{ g}$
(c) $3.15 \times 10^{-5} \text{ g}$
(d) $6.3 \times 10^{-5} \text{ g}$

66. Electrolysis of a solution of MnSO_4 in aqueous sulphuric acid is a method for the preparation of MnO_2 as per the reaction: $\text{Mn}^{2+}(\text{aq}) + 2\text{H}_2\text{O} \rightarrow \text{MnO}_2(\text{s}) + 2\text{H}^+(\text{aq}) + \text{H}_2(\text{g})$. Passing a current of 19.3 A for 2 h gives only 52.2 g of MnO_2 . The current efficiency is (Mn = 55)

- (a) 8.33%
(b) 83.33%
(c) 41.67%
(d) 100%

67. Copper sulphate solution (250 ml) was electrolysed using a platinum anode and a copper cathode. A constant current of 2 mA was passed for 19.3 min. It was found that after electrolysis the absorbance of the solution was reduced to 50% of its original value. Calculate the concentration of copper sulphate in the solution to begin with.
- (a) 9.6×10^{-5} M
(b) 4.8×10^{-5} M
(c) 2.4×10^{-5} M
(d) 1.2×10^{-5} M
68. Calculate the quantity of electricity that would be required to reduce 12.3 g of nitrobenzene to aniline if the current efficiency for the process is 50%. If the potential drop across the cell is 3.0 V, how much energy will be consumed?
- (a) 0.3 F, 86.85 kJ
(b) 0.6 F, 173.7 kJ
(c) 1.2 F, 347.4 kJ
(d) 0.6 F, 173.7 kJ
69. Perdisulphuric acid, $\text{H}_2\text{S}_2\text{O}_8$ can be prepared by electrolytic oxidation of H_2SO_4 , oxygen and hydrogen gases are by products. In such an electrolysis, 4.54 L of H_2 and 2.27 L of O_2 were generated at STP. What is the mass of $\text{H}_2\text{S}_2\text{O}_8$ formed?
- (a) 0
(b) 77.6 g
(c) 38.8 g
(d) 19.4 g
70. During the discharge of a lead storage battery, the density of sulphuric acid fell from 1.5 to 1.1 g/ml. Sulphuric acid of density 1.5 g/ml is 40% H_2SO_4 , by weight, and that of density 1.1 g/ml is 10% H_2SO_4 , by weight. The battery holds 3.6 L of the acid and the volume remained practically constant during the discharge. Calculate the number of ampere-hours which the battery should have been used. The electrode reactions are:
- $$\text{Pb} + \text{SO}_4^{2-} \rightarrow \text{PbSO}_4 + 2\text{e}^-$$
- $$\text{PbO}_2 + 4\text{H}^+ + \text{SO}_4^{2-} + 2\text{e}^- \rightarrow \text{PbSO}_4 + 2\text{H}_2\text{O}$$
- (a) 965
(c) 1930
- (b) 482.5
(d) 241.25
71. A cell whose resistance, when filled with 0.1 M – KCl is 200 Ω , is measured to be 6400 Ω , when filled with 0.003 M – NaCl solution. What is the molar conductance of NaCl solution, in $\Omega^{-1}\text{cm}^2\text{mol}^{-1}$ if the molar conductance of 0.1 M – KCl is 120 $\Omega^{-1}\text{cm}^2\text{eq}^{-1}$?
- (a) 41.67
(c) 125
- (b) 250
(d) 375
72. At 18°C, the ionic mobilities of NH_4^+ and ClO_4^- ions are 6.6×10^{-8} and 5.4×10^{-8} $\text{m}^2\text{volt}^{-1}\text{s}^{-1}$ at infinite dilution. What is the molar conductance of ammonium chromate solution in $\Omega^{-1}\text{m}^2\text{mol}^{-1}$?
- (a) 0.01795
(c) 1.2×10^{-8}
- (b) 0.01158
(d) 1.86×10^{-7}
73. At 25°C, the molar conductance at infinite dilution for HCl solution is $4.25 \times 10^{-2}\Omega^{-1}\text{m}^2\text{mol}^{-1}$, while its specific conductance is $382.5\Omega^{-1}\text{m}^{-1}$. If the degree of dissociation is 90%, the molarity of solution is
- (a) 0.9 M
(b) 1.0 M
(c) 0.1 M
(d) 1.1 M
74. A big irregular shaped vessel contained water, conductivity of which was $2.56 \times 10^{-3}\text{S}^{-1}\text{m}^{-1}$. 585 g of NaCl was then added to the water and conductivity after the addition of NaCl, was found to be $3.06 \times 10^{-3}\text{S}^{-1}\text{m}^{-1}$. The molar conductivity of NaCl at this concentration is $1.5 \times 10^{-2}\text{S}^{-1}\text{m}^2\text{mol}^{-1}$. The capacity of vessel if it is fulfilled with water, is
- (a) 3×10^4 l
(c) 3×10^8 l
- (b) 30 l
(d) 3×10^5 l

75. The molar conductivity of 0.10 M solution of MgCl_2 is $100 \text{ mho cm}^2 \text{ mol}^{-1}$, at 25°C . A cell with electrodes that are 1.50 cm^2 in surface area and 0.50 cm apart is filled with 0.10 M MgCl_2 solution. How much current will flow when the potential difference between the electrodes is 5 volts?
- (a) 0.03 A (b) 3.0 A
(c) 0.15 A (d) 15 A
76. For Na^+ , the value of symbol λ_m° is $50.0 \Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$. The speed of Na^+ ion in the solution, if in the cell, electrodes are 5 cm apart and to which a potential of 19.3 volt is applied is
- (a) $2 \times 10^{-3} \text{ cm/s}$
(b) $1 \times 10^{-3} \text{ cm/s}$
(c) $2 \times 10^{-4} \text{ cm/s}$
(d) $2 \times 10^{-2} \text{ cm/s}$
77. The conductivity of a saturated solution of AgCl at 298 K was found to be $3.40 \times 10^{-6} \Omega^{-1} \text{ cm}^{-1}$; the conductivity of water used to make up the solution was $1.60 \times 10^{-6} \Omega^{-1} \text{ cm}^{-1}$. Determine the solubility of AgCl in water in mole per litre at 298 K. The equivalent conductivity of AgCl at infinite dilution is $150.0 \Omega^{-1} \text{ cm}^2 \text{ eq}^{-1}$.
- (a) 1.44×10^{-10}
(b) 1.2×10^{-5}
(c) 3.33×10^{-5}
(d) 1.2×10^{-8}
78. Resistance of 0.2 M solution of an electrolyte is 50Ω . The specific conductance of the solution is 1.4 S m^{-1} . The resistance of 0.5 M solution of the same electrolyte is 280Ω . The molar conductivity of 0.5 M solution of the electrolyte (in $\text{S m}^2 \text{ mol}^{-1}$) is
- (a) 5×10^{-3} (b) 5×10^3
(c) 5×10^2 (d) 5×10^{-4}
79. The conductivity of a saturated solution containing AgA ($K_{\text{sp}} = 3 \times 10^{-14}$) and AgB ($K_{\text{sp}} = 1 \times 10^{-14}$) is $3.75 \times 10^{-8} \Omega^{-1} \text{ cm}^{-1}$. If the limiting molar conductivity of Ag^+ and A^- ion is 60 and $80 \Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$, respectively, the limiting molar conductivity of B^- (in $\Omega^{-1} \text{ cm}^2 \text{ mol}^{-1}$) is
- (a) 135
(b) 67.5
(c) 270
(d) 190
80. The conductivity of saturated solution of $\text{Ba}_3(\text{PO}_4)_2$ is $1.2 \times 10^{-5} \Omega^{-1} \text{ cm}^{-1}$. The limiting equivalent conductivities of BaCl_2 , K_3PO_4 and KCl are 160, 140 and $100 \Omega^{-1} \text{ cm}^2 \text{ eq}^{-1}$, respectively. The solubility product of $\text{Ba}_3(\text{PO}_4)_2$ is
- (a) 10^{-5}
(b) 1.08×10^{-23}
(c) 1.08×10^{-25}
(d) 1.08×10^{-27}

Section B (One or More than one Correct)

- Which of the following statement(s) differentiate between electrochemical cell and electrolytic cell?
 - Spontaneous or non-spontaneous nature of the chemical process.
 - Chemical reactions occurring at the electrodes.
 - Positive and negative nature of anode.
 - Dependence on Faraday's law.
- Pick up the false statement(s):
 - The net chemical change in a galvanic cell reaction is always redox reactions.
 - In a galvanic cell made of cobalt and cadmium electrodes, cobalt electrode acts as anode.
 - Standard potential increases with increasing concentration of the electrolyte.
 - Calomel electrode is a reference electrode having 0.00 volt potential.

3. Consider the cell: $\text{Ag(s)}, \text{AgCl(s)} \mid \text{KCl (0.1 M)} \mid \text{Hg}_2\text{Cl}_2\text{(s)}, \text{Hg(l)}$. The cell potential
 - (a) increases on increasing concentration of Cl^- ions.
 - (b) decreases on decreasing concentration of Cl^- ions.
 - (c) is independent of concentration of Cl^- ions.
 - (d) is independent of amounts of AgCl and Hg_2Cl_2 .
4. The passage of electricity in the Daniel cell when Zn and Cu electrodes are connected
 - (a) from Cu to Zn inside the cell
 - (b) from Cu to Zn outside the cell
 - (c) from Zn to Cu inside the cell
 - (d) from Zn to Cu outside the cell
5. From an electrolyte, one mole of electron will deposit at cathode
 - (a) 63.5 g of Cu
 - (b) 24 g of Mg
 - (c) 11.5 g of Na
 - (d) 9.0 g of Al
6. Which of the following statements is/are correct?
 - (a) The conductivity of molten NaCl is due to movement of Na^+ and Cl^- ions.
 - (b) Solid NaCl is good conductor of electricity.
 - (c) Molten sodium is a good conductor because of mobile electrons.
 - (d) Resistivity is reciprocal of molar conductivity of electrolyte.
7. Ionic conductance at infinite dilution of Al^{3+} and SO_4^{2-} ions are 60 and 80 $\Omega^{-1} \text{cm}^2 \text{eq}^{-1}$, respectively. The correct detail(s) regarding $\text{Al}_2(\text{SO}_4)_3$ is/are
 - (a) the molar conductance is 140 $\Omega^{-1} \text{cm}^2 \text{mol}^{-1}$.
 - (b) the equivalent conductance is 140 $\Omega^{-1} \text{cm}^2 \text{eq}^{-1}$.
 - (c) the molar conductance is 840 $\Omega^{-1} \text{cm}^2 \text{mol}^{-1}$.
 - (d) the molar conductance is 23.33 $\Omega^{-1} \text{cm}^2 \text{mol}^{-1}$.
8. The function(s) of salt bridge in a cell is/are
 - (a) It maintains standard electrode potential of cell constant which depends on several factors.
 - (b) It completes the electrical circuit.
 - (c) It departs both the solutions from each other.
 - (d) It maintains the electrical neutrality of both electrolytic solutions.
9. The reactions taking place in the dry cell are:

$$\text{Anode: Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$$

$$\text{Cathode: } 2\text{MnO}_2 + 2\text{NH}_4^+ + 2\text{e}^- \rightarrow \text{Mn}_2\text{O}_3 + 2\text{NH}_3 + \text{H}_2\text{O}$$

The minimum mass of reactants, if a dry cell is to generate 0.25A for 9.65 h, are ($\text{Mn} = 55$, $\text{Zn} = 65.4$) (neglect any other chemical reactions occurring in the cell)

 - (a) 2.943 g Zn
 - (b) 7.83 g MnO_2
 - (c) 1.62 g NH_4^+
 - (d) 3.915 g MnO_2
10. EMF of the cell: $\text{Cd(s)} \mid \text{CdCl}_2 \cdot 5\text{H}_2\text{O (sat.)} \mid \text{AgCl(s)} \mid \text{Ag(s)}$ is +0.70 V at 0°C and +0.60 V at 50°C . If ΔH° and ΔS° are temperature independent, then the correct information(s) regarding the cell reaction is/are
 - (a) $\Delta G^\circ = -115.8 \text{ kJ at } 50^\circ\text{C}$
 - (b) $\Delta G^\circ = 135.1 \text{ kJ at } 0^\circ\text{C}$
 - (c) $\Delta S^\circ = -386 \text{ J/K}$
 - (d) $\Delta H^\circ = -221.178 \text{ kJ}$
11. Two litre solution of a buffer mixture containing 1.0 M – NaH_2PO_4 and 1.0 M – Na_2HPO_4 is placed in two compartments (one litre in each) of an electrolytic cell. The platinum electrodes are inserted in each compartment and 1.25 A current is passed for 965 min. Assuming electrolysis of only water at each compartment, what will be pH in each compartment after passage of above charge? (pK_a for $\text{H}_2\text{PO}_4^- = 2.15$, $\log 7 = 0.85$)
 - (a) Anode: 3.00
 - (b) Cathode: 3.00
 - (c) Anode: 1.30
 - (d) Cathode: 1.30

12. For the reduction of NO_3^- ion in an aqueous solution, E° is +0.96 V. Values of E° for some metal ions are given below:
- $$\text{V}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{V}(\text{s}) \quad E^\circ = -1.19 \text{ V}$$
- $$\text{Fe}^{3+}(\text{aq}) + 3\text{e}^- \rightarrow \text{Fe}(\text{s}) \quad E^\circ = -0.04 \text{ V}$$
- $$\text{Au}^{3+}(\text{aq}) + 3\text{e}^- \rightarrow \text{Au}(\text{s}) \quad E^\circ = +1.40 \text{ V}$$
- $$\text{Hg}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Hg}(\text{l}) \quad E^\circ = +0.86 \text{ V}$$
- The pair(s) of metal that is(are) oxidized by NO_3^- in aqueous solution is(are)
- V and Hg
 - Hg and Fe
 - Fe and Au
 - Fe and V
13. Among the following, the intensive property is (properties are)
- Molar conductivity
 - Electromotive force
 - Resistance
 - Heat capacity
14. In a galvanic cell, the salt bridge
- does not participate chemically in the cell reaction.
 - stops the diffusion of ions from one electrode to another.
 - is necessary for the occurrence of the cell reaction.
 - ensures mixing of the two electrolytic solutions.
15. A lead storage cell is discharged which causes the H_2SO_4 electrolyte to change from a concentration of 40% by weight (density = 1.260 g/ml) to 28%, by weight. The original volume of electrolyte was 1 L. Identify the correct statement(s):
- The overall cell reaction is: $\text{Pb}(\text{s}) + \text{PbO}_2(\text{s}) + 2\text{H}_2\text{SO}_4(\text{aq}) \rightarrow 2\text{PbSO}_4(\text{s}) + 2\text{H}_2\text{O}(\text{l})$.
 - A total of 2.0 moles of H_2SO_4 is reacted.
 - The total charge released from anode of the cell is 1.93×10^5 coulomb.
 - The mass of electrolytic solution has decreased.

Section C (Comprehensions)

Comprehension I

Suppose that the S.H.E. was arbitrarily assigned a value of 1.00 V for $2\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$. What would this do to the observed voltage under standard condition for each of the following: (Given: $E^\circ_{\text{Zn}^{2+}|\text{Zn}} = -0.76 \text{ V}$, $E^\circ_{\text{Cu}^{2+}|\text{Cu}} = +0.34 \text{ V}$ when S.H.E. is 0.00 V)

- $E^\circ_{\text{Zn}^{2+}|\text{Zn}}$
 - +0.34 V
 - 0.76 V
 - +1.76 V
 - 0.34 V
- $E^\circ_{\text{Cu}^{2+}|\text{Cu}}$
 - 0.76 V
 - +1.34 V
 - +1.76 V
 - +0.34 V
- Zn–Cu Cell
 - +1.10 V
 - 1.10 V
 - 0.0 V
 - Indeterminate

Comprehension II

The standard reduction potential of the $\text{Ag}^+|\text{Ag}$ electrode at 298 K is 0.80 V. The solubility product of AgI is 6.4×10^{-17} at 298 K. ($2.303RT/F = 0.06$)

- The potential of $\text{Ag}^+|\text{Ag}$ electrode in a saturated solution of AgI at 298 K is
 - 0.314 V
 - +0.314 V
 - 0.172 V
 - +0.172 V
- The standard reduction potential of $\text{I}^-|\text{AgI}|\text{Ag}$ electrode at 298 K is
 - 0.314 V
 - +0.314 V
 - 0.172 V
 - +0.172 V
- The potential of $\text{I}^- (0.04 \text{ M})|\text{AgI}|\text{Ag}$ electrode at 198 K is
 - 0.088 V
 - +0.088 V
 - 0.172 V
 - +0.172 V

Comprehension III

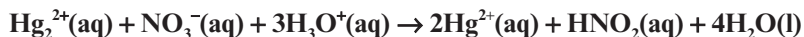
The Edison storage cell is represented as $\text{Fe(s)} | \text{FeO(s)} | \text{KOH(aq)} | \text{Ni}_2\text{O}_3\text{(s)} | \text{Ni(s)}$. The half-cell reactions are



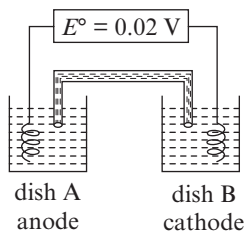
- What is the cell reaction?
 - $\text{Ni}_2\text{O}_3\text{(s)} + \text{Fe(s)} \rightarrow 2\text{NiO(s)} + \text{FeO(s)}$
 - $2\text{NiO(s)} + \text{FeO(s)} \rightarrow \text{Ni}_2\text{O}_3\text{(s)} + \text{Fe(s)}$
 - $\text{Ni}_2\text{O}_3\text{(s)} + \text{FeO(s)} \rightarrow 2\text{NiO(s)} + \text{Fe(s)} + \text{O}_2$
 - None of these
- What is the standard cell EMF?
 - 1.27 V
 - 0.47 V
 - 1.27 V
 - 0.47 V
- How does cell EMF depend on increasing the concentration of KOH?
 - increases
 - decreases
 - remains unaffected
 - none of these
- What is the maximum amount of electrical energy that can be obtained from one mole of Ni_2O_3 ?
 - 2.54 J
 - 245.11 kJ
 - 122.56 kJ
 - 61.28 kJ

Comprehension IV

The cell potential for the unbalanced chemical reaction:



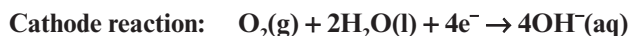
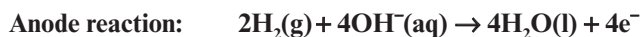
is measured under standard conditions in the electrochemical cell shown in the accompanying diagram.



11. In which dish, the solution is acidic?
 (a) Dish A (b) Dish B
 (c) Both (d) None
12. How many moles of electrons pass through the circuit when 0.60 mole of Hg^{2+} and 0.30 mole of HNO_2 are produced in the cell that contains 0.50 mole of Hg_2^{2+} and 0.40 mole of NO_3^- at the beginning of the reaction?
 (a) 0.30 (b) 0.60
 (c) 0.15 (d) 1.20
13. How long will it take to produce 0.10 mole of HNO_2 by this reaction if a current of 10 A passes through the cell?
 (a) 965 s
 (b) 193 s
 (c) 1930 s
 (d) 482.5 s
-

Comprehension V

A fuel cell is the device to convert the energy of a fuel into electrical energy without the use of heat engine, where the fuel is burnt directly. Such conversions are possible because the combustion reactions are essentially redox reactions and highly exothermic as well as highly exergonic. Electrical energy can be obtained indefinitely from a fuel cell as long as the outside supply of fuel is maintained. In hydrogen–oxygen fuel cell, the following reactions take place:

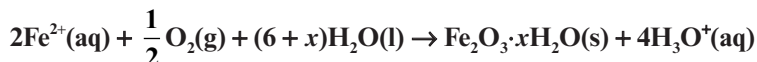


The overall reaction has a value of $\Delta H^\circ = -285.8 \text{ kJ/mol}$ and $\Delta G^\circ = -237.39 \text{ kJ/mol}$ at 25°C .

14. What is the standard EMF of the cell?
 (a) 0.615 V (b) 1.23 V
 (c) 2.46 V (d) 0.74 V
15. What volume of gaseous H_2 (at STP), when combined with excess O_2 in the fuel cell, is needed to produce 23.739 kJ of useful work under standard conditions?
 (a) 2.27 L (b) 4.54 L
 (c) 1.13 L (d) 2.24 L
16. Suppose the concentration of hydroxide ion in the cell is doubled at 298 K. The cell voltage will be
 (a) reduced by half
 (b) increased by a factor of 2
 (c) increased by a factor of 4
 (d) unchanged
17. What is the approximate value of ΔS° for the fuel cell reaction at 25°C ?
 (a) -0.1624 JK^{-1}
 (b) -162.4 JK^{-1}
 (c) $+162.4 \text{ JK}^{-1}$
 (d) $+0.1624 \text{ JK}^{-1}$
18. The theoretical efficiency of the fuel cell is given by
 (a) 83.06%
 (b) 100%
 (c) 67.53%
 (d) 97.88%
-

Comprehension VI

Billions of dollars are spent each year to replace or prevent the corrosion and subsequent destruction of government property such as ships, bridges, and metallic piping. Corrosion may result in decreased structural integrity and eventual ineffectuality. For example, a naval ship made from iron is highly susceptible to corrosion. One part of the hull acts as the anode, while another acts as the cathode, the iron hull itself connects the two parts, completing the circuit. Part of the metal hull begins oxidizing to Fe^{2+} in the presence of H_2O and O_2 , the reduction reaction proceeds with an $E_{\text{red}} = 1.229 \text{ V}$. It is there that Fe^{2+} ions migrate from the anode and are further oxidized as



The Fe_2O_3 , or rust, formed is only a small part of the problem. Pitting, or loss of solid metal, weakens the structure of the hull and may lead to serious damage and destruction of the hull. Using a sacrificial anode often successfully prevents corrosion. This metal, more easily oxidized than iron. Although the sacrificial anode eventually corrodes as well, the cost of replacing it is far less than that of replacing the iron hull.

Half-reaction	Potential
$\text{O}_2 + 4\text{H}^+ + 4\text{e}^- \rightarrow 2\text{H}_2\text{O}$	$E^\circ = 1.229 \text{ V}$
$\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}^{2+}$	$E^\circ = 0.771 \text{ V}$
$\text{Cu}^+ + \text{e}^- \rightarrow \text{Cu}$	$E^\circ = 0.521 \text{ V}$
$\text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^- \rightarrow 4\text{OH}^-$	$E^\circ = 0.401 \text{ V}$
$\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$	$E^\circ = 0.342 \text{ V}$
$\text{Pb}^{2+} + 2\text{e}^- \rightarrow \text{Pb}$	$E^\circ = -0.126 \text{ V}$
$\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$	$E^\circ = -0.447 \text{ V}$
$\text{Mn}^{2+} + 2\text{e}^- \rightarrow \text{Mn}$	$E^\circ = -1.185 \text{ V}$
$\text{Mg}^{2+} + 2\text{e}^- \rightarrow \text{Mg}$	$E^\circ = -2.372 \text{ V}$

19. Which of the following reactions may occur at the anode of the iron hull?
 - (a) $\text{Fe}^{2+}(\text{aq}) \rightarrow \text{Fe}^{3+}(\text{aq}) + \text{e}^-$
 - (b) $\text{Fe}(\text{s}) \rightarrow \text{Fe}^{2+}(\text{aq}) + 2\text{e}^-$
 - (c) $\text{O}_2(\text{g}) + 4\text{H}_3\text{O}^+(\text{aq}) \rightarrow 6\text{H}_2\text{O}(\text{l}) + 4\text{e}^-$
 - (d) $2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{OH}^-(\text{aq}) + \text{H}_2(\text{g}) + 2\text{e}^-$
20. What is the standard voltage generated by the corrosion reaction on the iron hull of a ship?
 - (a) -1.676 V
 - (b) -0.782 V
 - (c) 1.782 V
 - (d) 1.676 V
21. Which of the following materials could not serve as a sacrificial anode for lead?
 - (a) Copper
 - (b) Iron
 - (c) Magnesium
 - (d) Manganese
22. An ammeter reading shows a current of 0.50 A running through the iron hull.
 - (a) How many grams of $\text{Fe}(\text{s})$ are lost in 1.0 h ?
 - (a) 0.261 g
 - (b) 0.522 g
 - (c) 1.044 g
 - (d) 0.783 g
23. A ship that has sunk to the bottom of the ocean may still exhibit corrosion. One difference in the corrosion reaction between ship above water and one that has sunk to the bottom of the ocean may be that
 - (a) the oxidation reaction of the sunken ship does not include H_2O as a reactant.
 - (b) the oxidation reaction of the sunken ship does not include O_2 as a reactant.
 - (c) the reduction reaction of the sunken ship does not include H_2O as a reactant.
 - (d) the reduction reaction of the sunken ship does not include O_2 as a reactant.

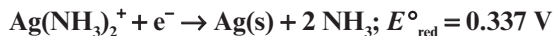
Comprehension VII

A current of 15.0 A is employed to plate nickel in a NiSO_4 solution. Both Ni and H_2 are formed at the cathode. The current efficiency with respect to formation of Ni is 60%. The density of nickel = 8.9 g/ml. (Ni = 58.7)

24. How much of nickel is plated on the cathode per hour?
 (a) 16.43 g (b) 32.85 g
 (c) 19.7 g (d) 9.85 g
25. What is the thickness of the plating if the cathode consists of a sheet of metal 4.0 cm^2 , which is to be coated on both sides?
 (a) 13.8 mm (b) 27.6 mm
 (c) 6.9 mm (d) 23.0 mm
26. What volume of H_2 at 0°C and 1 atm is formed per hour?
 (a) 6.27 l (b) 3.76 l
 (c) 2.5 l (d) 5.01 l
27. At the end of the electrolysis, how many grams of the gaseous product appear at the anode?
 (a) 4.48 g (b) 1.79 g
 (c) 2.69 g (d) 7.46 g

Comprehension VIII

Tollen's reagent is used for the detection of aldehyde when a solution of AgNO_3 is added to glucose with NH_4OH then gluconic acid is formed.



[Use $2.303RT/F = 0.0592$ and $FRT = 38.92$ at 298 K]

28. $2\text{Ag}^+ + \text{C}_6\text{H}_{12}\text{O}_6 + \text{H}_2\text{O} \rightarrow 2\text{Ag(s)} + \text{C}_6\text{H}_{12}\text{O}_7 + 2\text{H}^+$. Find $\ln K$ of this reaction.
 (a) 66.13 (b) 58.38
 (c) 28.30 (d) 46.29
29. When ammonia is added to the solution, pH is raised to 1 L. Which half-cell reaction is affected by pH and by how much?
 (a) E_{oxd} will increase by a factor of 0.65 from E°_{oxd}
 (b) E_{oxd} will decrease by a factor of 0.65 from E°_{oxd}
 (c) E_{red} will increase by a factor of 0.65 from E°_{red}
 (d) E_{red} will decrease by a factor of 0.65 from E°_{red}
30. Ammonia is always added in this reaction. Which of the following must be incorrect?
 (a) NH_3 combines with Ag^+ to form a complex
 (b) $\text{Ag}(\text{NH}_3)_2^+$ is a stronger oxidizing agent than Ag^+
 (c) In absence of ammonia, silver salt of gluconic acid is formed
 (d) NH_3 has affected the standard reduction potential of glucose–gluconic acid electrode

Comprehension IX

The standard potential of the following cell is 0.23 V at 15°C and 0.21 V at 35°C.



Given: The standard reduction potential of the $\text{Ag}^+(\text{aq}) \mid \text{Ag(s)}$ couple is 0.80 V at 25°C.

ΔH° and ΔS° for the cell reaction remain unchanged in the range 15°C to 35°C.

31. The overall cell reaction is
- $\text{H}_2(\text{g}) + 2\text{AgCl(s)} \rightarrow 2\text{HCl(aq)} + 2\text{Ag(s)}$
 - $\text{H}_2(\text{g}) + 2\text{Ag}^+(\text{aq}) \rightarrow 2\text{H}^+(\text{aq}) + 2\text{Ag(s)}$
 - $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl(s)}$
 - $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2\text{HCl(aq)}$
32. ΔH° for the cell reaction per mol of AgCl is
- 99,974 J
 - 49,987 J
 - +99,974 J
 - +49,987 J
33. ΔS° for the cell reaction per mol of AgCl is
- 193 J/K
 - 96.5 J/K
 - +96.5 J/K
 - +193 J/K
34. The solubility of AgCl in water at 25°C is ($2.303RT/F = 0.058$)
- 10^{-10} M
 - 10^{-12} M
 - 10^{-5} M
 - 10^{-6} M

Comprehension X

For the reaction: $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightleftharpoons \text{AgCl(s)}$. Given

Species	ΔG_f° (kJ / mol) at 25°C
$\text{Ag}^+(\text{aq})$	+77
$\text{Cl}^-(\text{aq})$	-129
AgCl(s)	-109



35. The cell representations of the above reaction and E_{cell}° are
- $\text{Ag(s)} \mid \text{AgCl(s)} \mid \text{Cl}^-(\text{aq}) \parallel \text{Ag}^+(\text{aq}) \mid \text{Ag(s)}; E_{\text{cell}}^\circ = 0.59 \text{ V}$
 - $\text{Ag(s)} \mid \text{Ag}^+(\text{aq}) \parallel \text{Cl}^-(\text{aq}) \mid \text{AgCl(s)} \mid \text{Ag(s)}; E_{\text{cell}}^\circ = 0.59 \text{ V}$
 - $\text{Ag(s)} \mid \text{Cl}^-(\text{aq}) \parallel \text{Ag}^+(\text{aq}) \mid \text{Ag}; E_{\text{cell}}^\circ = 0.59 \text{ V}$
 - $\text{Ag(s)} \mid \text{Ag}^+(\text{aq}) \parallel \text{Cl}^-(\text{aq}) \mid \text{AgCl(s)} \mid \text{Ag(s)}; E_{\text{cell}}^\circ = 1.18 \text{ V}$
36. The solubility product of AgCl at 298 K is
- 2×10^{-10}
 - 10^{-10}
 - e^{-10}
 - 10^{-5}
37. A quantity of 6.539×10^{-2} g of metallic zinc is added to 100 ml of saturated solution of AgCl. The value of $\log \frac{[\text{Zn}^{2+}]}{[\text{Ag}^+]^2}$ is (Zn = 65.39)
- 5.288
 - 52.88
 - 528.8
 - 26.44
38. How many moles of Ag will be precipitated in the above reaction?
- 10^{-3}
 - 2×10^{-3}
 - 10^{-6}
 - 10^{-5}

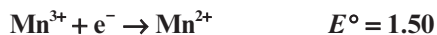
Comprehension XI

A sample of water from a large swimming pool has a resistance of $10,000\ \Omega$ at 25°C , when placed in a conductivity cell. When filled with $0.02\ \text{M}$ – KCl solution, the cell has a resistance of $100\ \Omega$ at 25°C . An amount of $585\ \text{g}$ of NaCl was dissolved in the pool, which was thoroughly stirred. A sample of this solution gave a resistance of $8000\ \Omega$. The molar conductivity of NaCl at this concentration is $125\ \Omega^{-1}\ \text{cm}^2\ \text{mol}^{-1}$ and molar conductivity of $0.02\ \text{M}$ – KCl is $200\ \Omega^{-1}\ \text{cm}^2\ \text{mol}^{-1}$.

39. Cell constant of the conductivity cell is
- $4\ \text{cm}^{-1}$
 - $0.4\ \text{cm}^{-1}$
 - $40\ \text{cm}^{-1}$
 - $0.04\ \text{cm}^{-1}$
40. Conductivity of water (in $\Omega^{-1}\ \text{cm}^{-1}$) is
- 0.4
 - 0.04
 - 0.0004
 - 0.00004
41. Volume (in L) of water in the pool is
- 1,25,000
 - 12,500
 - 1250
 - 125

Comprehension XII

Redox reactions play a pivotal role in chemistry and biology. The values of standard redox potential (E°) of two half-cell reactions decide which way the reaction is expected to proceed. A simple example is Daniel cell in which zinc goes into solution and copper gets deposited. Given below are a set of half-cell reactions (acidic medium) along with their E° (V with respect to normal hydrogen electrode) values. Using this data, obtain the correct explanations to questions below:



42. Among the following, identify the correct statement?
- Chloride ion is oxidized by O_2
 - Fe^{2+} is oxidized by iodine
 - Iodide ion is oxidized by chlorine
 - Mn^{2+} is oxidized by chlorine
43. While Fe^{3+} is stable, Mn^{3+} is not stable in acid solution because
- O_2 oxidizes Mn^{2+} to Mn^{3+}
 - O_2 oxidizes both Mn^{2+} to Mn^{3+} and Fe^{2+} to Fe^{3+}
 - Fe^{3+} oxidizes H_2O to O_2
 - Mn^{3+} oxidizes H_2O to O_2

Comprehension XIII

The concentration of potassium ions inside a biological cell is at least 20 times higher than the outside. The resulting potential difference across the cell is important in several processes such as transmission of nerve impulses and maintaining the ion balance. A simple model for such a concentration cell involving a metal 'M' is



For the above electrolytic cell, the magnitude of the cell potential, $|E_{\text{cell}}| = 70 \text{ mV}$.

44. For the above cell
- $E_{\text{cell}} < 0$; $\Delta G > 0$
 - $E_{\text{cell}} > 0$; $\Delta G < 0$
 - $E_{\text{cell}} < 0$; $\Delta G^\circ > 0$
 - $E_{\text{cell}} > 0$; $\Delta G^\circ < 0$
45. If the 0.05 molar solution of M^+ is replaced by a 0.0025 molar M^+ solution, then the magnitude of the cell potential would be
- 35 mV
 - 70 mV
 - 140 mV
 - 700 mV
-

Comprehension XIV

The electrochemical cell shown below is a concentration cell.



The EMF of the cell depends on the difference in concentrations of M^{2+} ions at the two electrodes. The EMF of the cell at 298 K is 0.059 V.

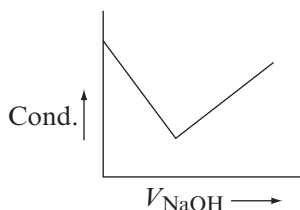
46. The value of ΔG (kJ mol^{-1}) for the given cell is (take $1 \text{ F} = 96500 \text{ C mol}^{-1}$)
- 5.7
 - 5.7
 - 11.4
 - 11.4
47. The solubility product (K_{sp} ; $\text{mol}^3 \text{ dm}^{-3}$) of MX_2 at 298 K based on the information available for the given concentration cell is (take $2.303 \times R \times 298/\text{F} = 0.059$)
- 1×10^{-15}
 - 4×10^{-15}
 - 1×10^{-12}
 - 4×10^{-12}
-

Comprehension XV

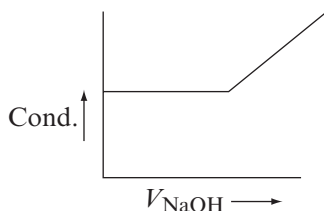
Conductance measurements are frequently employed to find the end points of acid–base and other titrations. The principle involved is that electrical conductance of a solution depends upon the number and mobility of ions. In the conductometric titrations, the conductance of the resulting solution is measured at different stages on adding some volume of the standard solution and then a graph is plotted from the experimental observations to get the end point. Conductometric titrations have several advantages. Coloured solutions, titration of weak acid and weak base, etc. cannot be titrated by normal methods, but can be titrated successfully by this method. Further, no special care is needed in titration because the end point is determined graphically.

48. Which of the following graph truly represents the titration of HCl solution against NaOH solution?

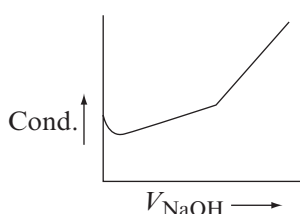
(a)



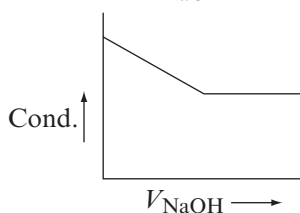
(b)



(c)

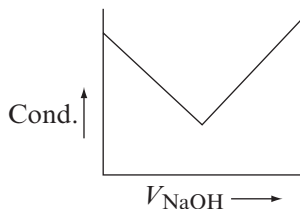


(d)

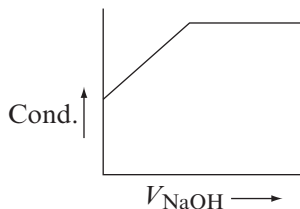


49. Which of the following graph truly represents the titration of CH_3COOH solution against NaOH solution?

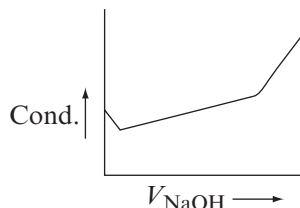
(a)



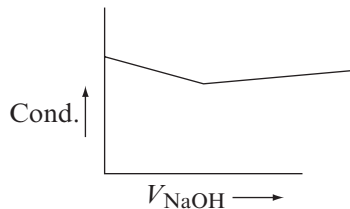
(b)



(c)

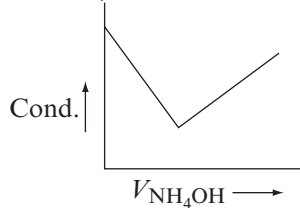


(d)

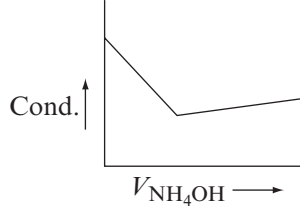


50. Which of the following graph truly represents the titration of HCl solution against NH_4OH solution?

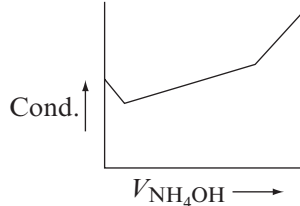
(a)



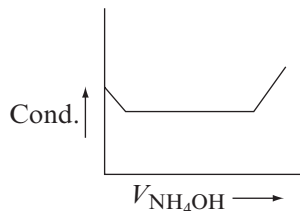
(b)



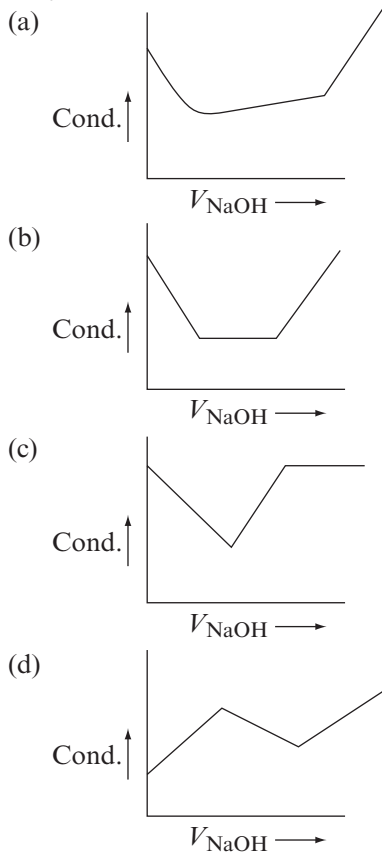
(c)



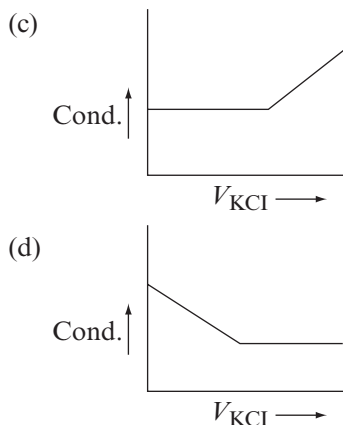
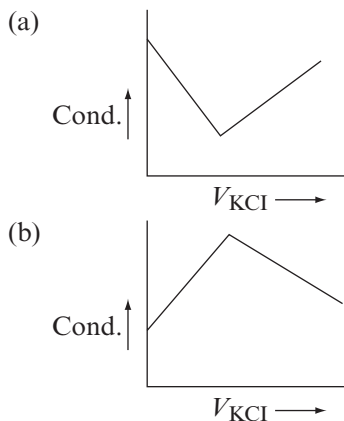
(d)



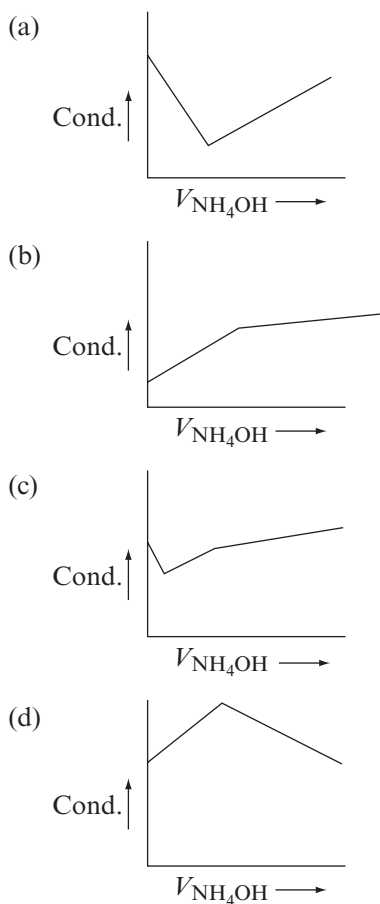
51. Which of the following graph truly represents the titration of a solution containing a mixture of HCl and CH_3COOH against NaOH solution?



52. Which of the following graph truly represents the titration of AgNO_3 solution against KCl solution?



53. Which of the following graph truly represents the titration of CH_3COOH solution against NH_4OH solution?



Section D (Assertion – Reason)

The following questions consist of two statements. Mark

- (a) If both statements are CORRECT, and **Statement II** is the CORRECT explanation of **Statement I**.
- (b) If both statements are CORRECT, and **Statement II** is NOT the CORRECT explanation of **Statement I**.
- (c) If **Statement I** is CORRECT, but **Statement II** is INCORRECT.
- (d) If **Statement I** is INCORRECT, but **Statement II** is CORRECT.

1. **Statement I:** Electrolysis of $\text{CuCl}_2(\text{aq})$ gives 1 mole of Cu and 1 mole of Cl_2 by the passage of suitable charge.

Statement II: Equal equivalents of Cu and Cl_2 are formed during the passage of same charge.

2. **Statement I:** Zinc metal can displace Ag metal from a solution containing the complex $[\text{Ag}(\text{CN})_2]^-$.

Statement II: $E^\circ_{\text{Zn}^{2+}, \text{Zn}}$ is greater than $E^\circ_{\text{Ag}^+, \text{Ag}}$.

3. **Statement I:** In electrolysis, the quantity of electricity needed for depositing 1 mole of silver is different from that required for 1 mole copper.

Statement II: The atomic masses of silver and copper are different.

4. **Statement II:** In an electrochemical cell, anode and cathode are, respectively, negative and positive electrode.

Statement II: At anode, oxidation takes place and at cathode reduction takes place.

5. **Statement I:** A metal having negative reduction potential when dipped in the solution of its own ions has a tendency to pass into the solution.

Statement II: Metals having negative reduction potential have large hydration energy.

6. **Statement I:** Specific conductance decreases with dilution while molar conductance increases.

Statement II: On dilution, number of ions per unit volume decreases but total number of ions increases considerably.

7. **Statement I:** When acidified zinc sulphate solution is electrolysed between zinc electrodes, it is zinc that is deposited at the cathode and evolution of hydrogen gas does not take place.

Statement II: The electrode potential of zinc becomes more negative than hydrogen as the overvoltage for the hydrogen evolution on zinc is quite large.

8. **Statement I:** The conductivity of solutions of different electrolytes in the same solvent and at a given temperature is same.

Statement II: The conductivity depends on the charge and size of the ions, the concentrations of ions and ease with which the ions move under potential gradient.

9. **Statement I:** Sodium ions are discharged in preference to hydrogen ions at a mercury cathode.

Statement II: Na^+ is a strong reducing agent in comparison to H^+ ion.

10. **Statement I:** The conductivity of metals decreases while that of electrolytic solution increases with increase in temperature.

Statement II: Electrons in metals are very tightly held by the nucleus and are not free to move.

11. **Statement I:** An electrochemical cell can be set up only when the redox reaction is spontaneous.

Statement II: A reaction is spontaneous if free energy change at constant temperature and pressure is negative.

12. **Statement I:** In the Daniel cell, if concentrations of Cu^{2+} and Zn^{2+} ions are doubled, the EMF of cell does not change.

Statement II: If the concentration of ions in contact with the metal is doubled, the electrode potential will be doubled.

13. **Statement I:** KCl, NaCl, NH_4Cl , etc., cannot be used in the salt bridge of a cell containing silver.

Statement II: A salt bridge contains concentrated solution of an inert electrolyte like KCl, KNO_3 , K_2SO_4 , etc., in agar-agar.

14. **Statement I:** During electrolysis of aqueous sodium acetate solution, the molar ratio of gases formed at cathode and anode is 1:3.

Statement II: Acetate ion discharges at anode and H^+ ion, at cathode.

15. **Statement I:** Electrode potential of any electrode will change on changing any of its intensive properties.

Statement II: Any intensive property of system changes on changing any of its properties, whether intensive or extensive.

Section E (Column Match)

1. Match Column I with Column II

Column I	Column II (Electrolysis product using inert electrodes)
(A) Dilute solution of HCl	(P) O_2 evolved at anode
(B) Dilute solution of NaCl	(Q) H_2 evolved at cathode
(C) Concentrated solution of NaCl	(R) Cl_2 evolved at anode
(D) AgNO_3 solution	(S) Ag deposited at cathode

2. An aqueous solution of 'X' is added slowly to an aqueous solution of 'Y' as shown in List I. The variation in conductivity of these reactions is given in List II. Match List I with List II.

List I X Y	List II
(A) $(\text{C}_2\text{H}_5)_3\text{N} + \text{CH}_3\text{COOH}$	(P) Conductivity decreases and then increases
(B) $\text{KI}(0.1 \text{ M}) + \text{AgNO}_3(0.01 \text{ M})$	(Q) Conductivity decreases and then does not change much
(C) $\text{CH}_3\text{COOH} + \text{KOH}$	(R) Conductivity increases and then does not change much
(D) $\text{NaOH} + \text{HI}$	(S) Conductivity does not change much and then increases

3. The standard reduction potential data at 25°C is given below:

$$E^\circ (\text{Fe}^{3+}, \text{Fe}^{2+}) = +0.77 \text{ V}$$

$$E^\circ (\text{Fe}^{2+}, \text{Fe}) = -0.44 \text{ V}$$

$$E^\circ (\text{Cu}^{2+}, \text{Cu}) = +0.34 \text{ V}$$

$$E^\circ (\text{Cu}^+, \text{Cu}) = +0.52 \text{ V}$$

$$E^\circ [\text{O}_2(\text{g}) + 4\text{H}^+ + 4\text{e}^- \rightarrow 2\text{H}_2\text{O}] = +1.23 \text{ V}$$

$$E^\circ [\text{O}_2(\text{g}) + 2\text{H}_2\text{O} + 4\text{e}^- \rightarrow 4\text{OH}^-] = +0.40 \text{ V}$$

$$E^\circ (\text{Cr}^{3+}, \text{Cr}) = -0.74 \text{ V}$$

$$E^\circ (\text{Cr}^{2+}, \text{Cr}) = -0.91 \text{ V}$$

Match E° of the redox pair in List I with the values given in List II.

List I	List II
(A) $E^\circ (\text{Fe}^{3+}, \text{Fe})$	(P) -0.18 V
(B) $E^\circ (4\text{H}_2\text{O} \rightarrow 4\text{H}^+ + 4\text{OH}^-)$	(Q) -0.4 V
(C) $E^\circ (\text{Cu}^{2+} + \text{Cu} \rightarrow 2\text{Cu}^+)$	(R) -0.04 V
(D) $E^\circ (\text{Cr}^{3+}, \text{Cr}^{2+})$	(S) -0.83 V

4. Match the following

Column I	Column II
(A) Concentration cell	(P) $\text{Ag} \text{AgCl} \text{Cl}^- \text{Ag}^+ \text{Ag}$
(B) Spontaneous cell reaction	(Q) $\text{Ag} \text{AgCl} \text{Cl}^- \text{Br}^- \text{AgBr} \text{Ag}$
(C) Non-spontaneous cell reaction	(R) $\text{Ag} \text{Ag}^+ (0.1 \text{ M}) \text{Ag}^+ (1.0 \text{ M}) \text{Ag}$
	(S) $\text{Ag} \text{AgCl} \text{Cl}^- (0.1 \text{ M}) \text{Cl}^- (1.0 \text{ M}) \text{AgCl} \text{Ag}$

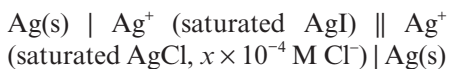
Section F (Subjective)

Single-digit Integer Type

- Percentage of aniline hydrochloride hydrolysed in its M/40 solution at 25°C is (Given: $E_{\text{C}_6\text{H}_5\text{NH}_2 \cdot \text{HCl} | \text{H}_2} = -0.18 \text{ V}$, $2.303RT/F = 0.06$)
- The standard reduction potential for $\text{Cu}^{2+} | \text{Cu}$ electrode is $+0.34 \text{ V}$. The solubility product of $\text{Cu}(\text{OH})_2$ is 1.0×10^{-19} . The pH of solution at which the reduction potential for the above electrode becomes 0.31 V , is ($2.303RT/F = 0.06$)
- The reduction potential at 25°C for $\text{Fe}^{3+} | \text{Fe}^{2+}$ electrode is $+0.718 \text{ V}$. If $E_{\text{Fe}^{2+} | \text{Fe}}^\circ = -0.44 \text{ V}$ and $E_{\text{Fe}^{3+} | \text{Fe}}^\circ = -0.04 \text{ V}$, the ratio of molar concentrations of Fe^{2+} to Fe^{3+} ions in solution is ($2.303RT/F = 0.06$, $\log 5 = 0.7$)

4. If $[\text{Fe}^{3+}]$ at equilibrium, when potassium iodide is added to a solution of Fe^{3+} initially at 0.50 M until $[\text{I}^-] = 1.0 \text{ M}$, is $x \times 10^{-5} \text{ M}$, the value of x is (Given $E^\circ_{\text{Fe}^{3+}|\text{Fe}^{2+}} = 0.77 \text{ V}$, $E^\circ_{\text{I}_2|\text{I}^-} = 0.53 \text{ V}$, $2.303RT/F = 0.06$)

5. If it is desired to construct the following galvanic cell:



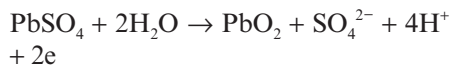
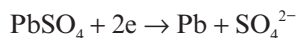
to have $E_{\text{cell}} = 0.102 \text{ V}$, what should be the value of x to get $[\text{Cl}^-]$, which must be present in the cathodic half-cell to achieve the desired EMF. Given K_{sp} of AgCl and AgI are 1.8×10^{-10} and 8.1×10^{-17} , respectively. ($2.303RT/F = 0.06$)

6. An alloy of lead (valency = 2)-thallium (valency = 1) containing 70% Pb and 30% Tl, by weight, can be electroplated onto a cathode from a perchloric acid solution. How many hours would be required to deposit 5.0 g of this alloy at a current of 1.10 A? (Pb = 208, Tl = 204)
7. Iridium was plated from a solution containing IrCl_6^{y-} for 2.0 h with a current of 0.075 A. The Iridium deposited on the cathode weighed 0.36 g. If the oxidation state of Ir in IrCl_6^{y-} is x , then the value of $(x + y)$ is (Ir = 192)
8. The electrolysis of cold sodium chloride solution produces sodium hypochlorite by reacting NaOH and Cl_2 thoroughly. How long (in days) will a cell operate to produce 10 L of 7.45% (by mass) solution of NaClO if the cell current is 2.5 A? Assume that the density of solution is 1.0 g/ml.
9. A volume of 500 ml of 0.1 M – CuSO_4 solution is electrolysed for 5 min at a current of 0.161 A. If Cu is produced at one electrode and oxygen at the other, the approximate pH of the final solution is

10. A test for complete removal of Cu^{2+} ions from a solution of Cu^{2+} is to add $\text{NH}_3(\text{aq})$. A blue colour signifies the formation of complex $[\text{Cu}(\text{NH}_3)_4]^{2+}$ having $K_f = 1.1 \times 10^{13}$ and thus confirms the presence of Cu^{2+} in solution. 250 ml of 0.1 M – CuSO_4 is electrolysed by passing a current of 5 A for 1351 s. After passage of this charge, sufficient quantity of NH_3 is added to electrolysed solution maintaining $[\text{NH}_3] = 0.10 \text{ M}$. If $[\text{Cu}(\text{NH}_3)_4]^{2+}$ is detectable up to its concentration as low as $1 \times 10^{-5} \text{ M}$, would a blue colour be shown by the electrolysed solution on addition of NH_3 . Mark '1', if the answer is 'yes' and mark '2', if the answer is 'no'.

11. By passing a certain amount of charge through NaCl solution, 9.08 L of chlorine gas were liberated at STP. When the same amount of charge is passed through a nitrate solution of metal M, 52.8 g of the metal was deposited. If the specific heat of metal is $0.032 \text{ Cal}^\circ\text{C}^{-1}\text{g}^{-1}$, the valency of metal is

12. The electrode reactions for charging of a lead storage battery are:



The electrolyte in the battery is an aqueous solution of sulphuric acid. Before charging, the specific gravity of the liquid was found to be 1.10 (16% H_2SO_4 by wt.). After charging for $\frac{965}{9}$ h, the specific gravity of the liquid was found to be 1.42 (40% H_2SO_4 by weight). If the battery contained 2 L of the liquid and the volume remains constant during charge, the average current (in A) used for charging the battery is

13. A dilute solution of KCl was placed between two platinum electrodes, 12 cm apart, across which a potential of 1.93 volts was applied. How far (in cm) would the K^+ ion move in 20 h at 25°C? Ionic conductance of K^+ ion at infinite dilution at 25°C is 7.5×10^{-3} mho m^2 mol^{-1} .
14. In the refining of silver by electrolytic method, what will be the mass of 72.8 g silver anode (60% pure, by weight), if 9.65 A current is passed for 1 h? ($Ag = 108$)
15. The pH of 100 L of concentrated KCl solution after the electrolysis for 10 s using 9.65 A at 298 K is

Four-digit Integer Type

1. The specific conductivity of a saturated solution of AgCl is 2.80×10^{-4} mho m^{-1} at 25°C. If $\lambda_{Ag^+}^\circ = 6.19 \times 10^{-3}$ mho m^2 mol^{-1} and $\lambda_{Cl^-}^\circ = 7.81$ mho m^2 mol^{-1} , the solubility of silver chloride (in order of 10^{-5} g l^{-1}) at 25°C, is
2. The electrode potential (in millivolts) of $2Ag(s) + S^{2-}(aq) \rightarrow Ag_2S(s) + 2e^-$ in a solution buffered at pH = 3 and which is also saturated with 0.1 M H_2S , is (Given: for H_2S , $K_{a1} = 10^{-8}$; $K_{a2} = 10^{-13}$, K_{sp} of $Ag_2S = 4 \times 10^{-48}$, $E_{Ag^+|Ag}^\circ = 0.80$ V, $\log 2 = 0.3$, $2.303RT/F = 0.06$)
3. An excess of liquid mercury is added to an acidified solution of 10^{-3} M Fe^{3+} . It is found that 10% of Fe^{3+} remains at equilibrium at 25°C. The value of $E_{Hg_2^{2+}|Hg}^\circ$ (in mV), assuming that the only reaction that occurs is: $2Hg + 2Fe^{3+} \rightarrow Hg_2^{2+} + 2Fe^{2+}$, is (Given: $E_{Fe^{3+}|Fe^{2+}}^\circ = 0.7724$ V, $\log 2 = 0.3$, $\log 3 = 0.48$, $2.303RT/F = 0.06$)
4. Determine potential (in mV) of the cell: Pt | Fe^{2+} , Fe^{3+} || $Cr_2O_7^{2-}$, Cr^{3+} , H^+ | Pt in which $[Fe^{2+}] = 0.75$ M, $[Fe^{3+}] = 0.75$ M, $[Cr_2O_7^{2-}] = 2$ M, $[Cr^{3+}] = 4$ M and $[H^+] = 1$ M. Given:
- $$Fe^{3+} + e^- \rightarrow Fe^{2+}; E^\circ = 0.77 \text{ V}$$
- $$14 H^+ + 6e^- + Cr_2O_7^{2-} \rightarrow 2Cr^{3+} + 7H_2O;$$
- $$E^\circ = 1.35 \text{ V}$$
- $$2.303RT/F = 0.06, \log 2 = 0.3, \log 3 = 0.48$$
5. An alloy weighing 2.70 mg of Pb–Ag was dissolved in desired amount of HNO_3 and volume was made 250 ml. A silver electrode was dipped in solution and E_{cell} of the cell:
- $$Pt, H_2(1 \text{ bar}) | H^+(1 \text{ M}) || Ag^+ | Ag$$
- was 0.50 V at 298 K. The percentage of lead in alloy is (Given: $E_{Ag^+|Ag}^\circ = 0.80$ V, $Ag = 108$, $2.303RT/F = 0.06$)
6. Given $E^\circ = 0.08$ V for $Fe^{3+}(\text{cyt b}) | Fe^{2+}(\text{cyt b})$ couple and $E^\circ = 0.20$ V for $Fe^{3+}(\text{cyt c}_1) | Fe^{2+}(\text{cyt c}_1)$ couple, where E° represents the standard state reduction potentials at pH = 7.0 at 25°C and cyt is an abbreviation for cytochromes. The value of K_{eq} for the reaction
- $$Fe^{3+}(\text{cyt c}_1) + Fe^{2+}(\text{cyt b}) \rightleftharpoons Fe^{2+}(\text{cyt c}_1) + Fe^{3+}(\text{cyt b})$$
- is ($2.303RT/F = 0.06$)
7. Calculate potential (in mV) of the cell: Cu | Mn(s) | $MnCl_2(0.001 \text{ M})$, $HCl(0.01 \text{ M})$ | $O_2(0.25 \text{ bar})$ | Pt | Cu. Given, $E^\circ = -1.185$ V for the $Mn^{2+} | Mn$ couple and 1.229 V for the $O_2 | H_2O, H^+$ couple. ($2.303RT/F = 0.06$, $\log 2 = 0.3$)
8. The temperature coefficient of the cell: Zn | $ZnCl_2$ | $AgCl(s)$ | Ag is +0.001 V/K. The entropy change (in J/K) accompanying this reaction, at 298 K is

9. For the reaction: $4\text{Al(s)} + 3\text{O}_2\text{(g)} + 6\text{H}_2\text{O(l)} + 4\text{OH}^-\text{(aq)} \rightleftharpoons 4\text{Al(OH)}_4^-\text{(aq)}$; $E^\circ_{\text{cell}} = 2.5\text{ V}$. If $\Delta_f G^\circ$ for $\text{H}_2\text{O(l)}$ and $\text{OH}^-\text{(aq)}$ are -280.0 and -156.25 kJ/mol , respectively, the magnitude of $\Delta_r G^\circ$ for $\text{Al(OH)}_4^-\text{(aq)}$ (in kJ/mol) is
10. In a Zn–MnO₂ Cell, the anode is made up of Zn and cathode of carbon rod surrounded by a mixture of MnO₂, Carbon, NH₄Cl and ZnCl₂ in aqueous base. If 8.7 g MnO₂ is present in cathodic compartment, how many days the dry cell will continue to give a current of $3.99 \times 10^{-3}\text{ A}$?
11. An amount of 19 g of molten SnCl₂ is electrolysed for some time. Inert electrodes are used. 1.19 g of tin is deposited at the cathode. No substance is lost during electrolysis. If the ratio of the masses of SnCl₂ and SnCl₄ after electrolysis is 261:x, the value of x is (Sn = 119)
12. An object whose surface area is 80 cm^2 is to be plated with an even layer of gold $8.0 \times 10^{-4}\text{ cm}$ thick. The density of gold is 19.7 g/cm^3 . The object is placed in a solution of $\text{Au(NO}_3)_3$ and a current of 2.4 A is applied. The time (in seconds) required for the electroplating to be completed, assuming that the layer of gold builds up evenly, is (Au = 197)
13. We have taken a saturated solution of AgBr. K_{sp} of AgBr is 12×10^{-14} . If 10^{-7} moles of AgNO₃ is added to 1 litre of this solution, find conductivity (specific conductance) of this solution in terms of $10^{-7}\text{ S}\cdot\text{m}^{-1}$ units. Given $\lambda^\circ(\text{Ag}^+) = 6 \times 10^{-3}\text{ S m}^2\text{ mol}^{-1}$, $\lambda^\circ(\text{Br}^-) = 8 \times 10^{-3}\text{ S m}^2\text{ mol}^{-1}$, $\lambda^\circ(\text{NO}_3^-) = 7 \times 10^{-3}\text{ S}\cdot\text{m}^2\text{ mol}^{-1}$.
14. On passing electricity through nitrobenzene solution, it is converted into azobenzene. The mass of azobenzene (in mg) produced, if the same quantity of electricity produces oxygen just sufficient to burn 96 g of fullerene (C₆₀), is
15. The conductivity of saturated solution of sparingly soluble salt, Ba₃(PO₄)₂, is $1.2 \times 10^{-5}\text{ ohm}^{-1}\text{cm}^{-1}$. The limiting equivalent conductances of BaCl₂, K₃PO₄ and KCl are 160, 140 and $100\text{ ohm}^{-1}\text{cm}^2\text{ eq}^{-1}$, respectively. The k_{sp} of Ba₃(PO₄)₂ (in the order of 10^{-25}) is

Answer Keys – Exercise II

Section A (Only one Correct)

- | | | | | | | | | | |
|---------|---------|---------|---------|---------|---------|---------|---------|---------|---------|
| 1. (c) | 2. (d) | 3. (b) | 4. (b) | 5. (a) | 6. (a) | 7. (d) | 8. (d) | 9. (b) | 10. (c) |
| 11. (c) | 12. (a) | 13. (a) | 14. (b) | 15. (b) | 16. (b) | 17. (a) | 18. (b) | 19. (d) | 20. (a) |
| 21. (b) | 22. (c) | 23. (a) | 24. (a) | 25. (b) | 26. (c) | 27. (d) | 28. (b) | 29. (b) | 30. (b) |
| 31. (b) | 32. (a) | 33. (c) | 34. (a) | 35. (b) | 36. (a) | 37. (d) | 38. (a) | 39. (a) | 40. (a) |
| 41. (a) | 42. (c) | 43. (a) | 44. (d) | 45. (a) | 46. (b) | 47. (c) | 48. (b) | 49. (b) | 50. (a) |
| 51. (b) | 52. (b) | 53. (a) | 54. (a) | 55. (a) | 56. (b) | 57. (d) | 58. (c) | 59. (a) | 60. (b) |
| 61. (c) | 62. (a) | 63. (a) | 64. (b) | 65. (d) | 66. (b) | 67. (a) | 68. (c) | 69. (c) | 70. (b) |
| 71. (c) | 72. (a) | 73. (c) | 74. (a) | 75. (c) | 76. (a) | 77. (b) | 78. (d) | 79. (c) | 80. (b) |

Section B (One or More than one Correct)

- | | | | |
|------------------|------------------------|------------------------|-------------------|
| 1. (a), (c) | 2. (c), (d) | 3. (c), (d) | 4. (b), (c) |
| 5. (d) | 6. (a), (c) | 7. (b), (c) | 8. (b), (c), (d) |
| 9. (a), (b), (c) | 10. (a), (b), (c), (d) | 11. (b), (c) | 12. (a), (b), (d) |
| 13. (a), (b) | 14. (a) | 15. (a), (b), (c), (d) | |

Section C

Comprehension I

1. (a) 2. (b) 3. (a)

Comprehension II

4. (b) 5. (d) 6. (b)

Comprehension III

7. (a) 8. (a) 9. (c) 10. (b)

Comprehension IV

11. (b) 12. (b) 13. (c)

Comprehension V

14. (b) 15. (a) 16. (d) 17. (b) 18. (a)

Comprehension VI

19. (b) 20. (d) 21. (a) 22. (b) 23. (d)

Comprehension VII

24. (d) 25. (a) 26. (c) 27. (a)

Comprehension VIII

28. (b) 29. (a) 30. (d)

Comprehension IX

31. (a) 32. (b) 33. (b) 34. (c)

Comprehension X

35. (a) 36. (b) 37. (b) 38. (c)

Comprehension XI

39. (b) 40. (d) 41. (a)

Comprehension XII

42. (c) 43. (d)

Comprehension XIII

44. (b) 45. (c)

Comprehension XIV

46. (d) 47. (b)

Comprehension XV

48. (a) 49. (c) 50. (b) 51. (a)
52. (c) 53. (b)

Section D (Assertion – Reason)

1. (a) 2. (c) 3. (b) 4. (a) 5. (b) 6. (c) 7. (a) 8. (d) 9. (c) 10. (c)
11. (b) 12. (c) 13. (b) 14. (b) 15. (c)

Section E (Column Match)

1. $A \rightarrow P, Q; B \rightarrow P, Q; C \rightarrow Q, R; D \rightarrow P, S$
2. $A \rightarrow R; B \rightarrow S; C \rightarrow Q; D \rightarrow P$
3. $A \rightarrow R; B \rightarrow S; C \rightarrow P; D \rightarrow Q$
4. $A \rightarrow R, S; B \rightarrow P, R; C \rightarrow Q, S$

Section F (Subjective)

Single-digit Integer Type

1. (4) 2. (9) 3. (5) 4. (5) 5. (4) 6. (1) 7. (0) 8. (4) 9. (2) 10. (1)
11. (3) 12. (2) 13. (9) 14. (8) 15. (9)

Four-digit Integer Type

1. (0287) 2. (0142) 3. (0815) 4. (0571) 5. (0090)
6. (0100) 7. (2453) 8. (0193) 9. (1300) 10. (0028)
11. (1520) 12. (0772) 13. (0055) 14. (0728) 15. (0108)