
DPP NO. 02
TOPIC : TYPES OF CHEMICAL REACTION

1. When lead nitrate reacts with potassium iodide, yellow precipitate of
(A) PbI_2 is formed (B) KNO_3 is formed
(C) $\text{Pb}(\text{NO}_3)_2$ is formed (D) PbIO_3 is formed
2. When ferrous hydroxide reacts with hydrochloric acid, and H_2O are produced
(A) FeCl_3 (B) FeCl_2 (C) FeCl_4 (D) FeCl
3. Which of the following reactions involves the combination of two elements ?
(A) $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$ (B) $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$
(C) $\text{SO}_2 + \frac{1}{2}\text{O}_2 \rightarrow \text{SO}_3$ (D) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
4. Which of the following is a displacement reaction ?
(A) $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca}(\text{OH})_2$ (B) $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
(C) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ (D) $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
5. $\text{AgNO}_{3(\text{aq})} + \text{NaCl}_{(\text{aq})} \rightarrow \text{AgCl}_{(\text{s})} + \text{NaNO}_{3(\text{aq})}$. Above reaction is
(A) Precipitation reaction (B) Double Displacement reaction
(C) Combination reaction (D) A and B both
6. Which of the following is a decomposition reaction ?
(A) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ (B) $\text{NH}_4\text{CNO} \rightarrow \text{H}_2\text{NCOONH}_2$
(C) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ (D) $\text{H}_2 + \text{I}_2 \rightarrow 2\text{HI}$
7. Which of the following is a displacement reaction ?
(A) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ (B) $\text{CaO} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O}$
(C) $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$ (D) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
8. Which of the following is not a decomposition reaction ?
(A) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ (B) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
(C) Digestion of food in body (D) $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
9. $\text{Fe}_2\text{O}_3 + 2\text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
The above reaction is an example of
(A) Combination reaction (B) Double displacement reaction
(C) Decomposition reaction (D) Displacement reaction
10. Neutralization reaction is an example of
(A) exothermic reaction (B) endothermic reaction (C) oxidation (D) none of these
11. What happens when $\text{CO}_2(\text{g})$ is bubbled through lime water
(i) in small amount ? (ii) in excess ?
12. Complete and balance the following reactions
(i) $\text{Pb}(\text{NO}_3)_2 \xrightarrow{\text{Heat}}$
(ii) $(\text{NH}_4)_2\text{Cr}_2\text{O}_7(\text{s}) \xrightarrow{\text{Heat}}$
(iii) $\text{NH}_3 + \text{CuO} \longrightarrow$
13. In the refining of silver, the recovery of silver from silver nitrate solution involved displacement by copper metal. Write down the reaction involved.
14. X, Y and Z are three elements which undergo chemical reactions according to following equations.
 $\text{X}_2\text{O}_3 + 2\text{Y} \rightarrow \text{Y}_2\text{O}_3 + 2\text{X}$; $3\text{ZSO}_4 + 2\text{Y} \rightarrow \text{Y}_2(\text{SO}_4)_3 + 3\text{Z}$; $3\text{ZO} + 2\text{X} \rightarrow \text{X}_2\text{O}_3 + 3\text{Z}$
Answer the following equations
(a) Which element is the most reactive ?
(b) Which element is the least reactive ?
15. (a) Why is combustion reaction an oxidation reaction ?
(b) How will you test whether the gas evolved in a reaction is hydrogen ?

Answers Key

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1.	A	2.	B	3.	B	4.	C	5.	D	6.	C	7.	C
8.	D	9.	D	10.	A								