## **ORGANIC CHEMISTRY**



## DPP No. 14

**Total Marks: 23** 

Max. Time: 23 min.

**Topic: Hydrocarbons** 

**Type of Questions** 

M.M., Min.

Single choice Objective ('-1' negative marking) Q.1 to Q.7

[21, 21]

True or False (no negative marking) Q.8

(3 marks, 3 min.) (2 marks, 2 min.)

[2, 2]

1.  $CH_3$ -CH-CH= $CH_2$   $\xrightarrow{H_3O^+}$  P (major), P is :

(A) CH<sub>3</sub>-CH-CH<sub>2</sub>-CH<sub>2</sub>-OH

(C) CH<sub>3</sub>-C-CH<sub>2</sub>-CH

2.  $CH=CH-CH_3 \xrightarrow{H_2O/H^*} P \text{ (major)}, P \text{ is :}$ 

3. which of the following alkene will give (P) on oxymercuration reduction reaction, (P) =  $\begin{array}{c} CH_3 \\ OH \end{array}$ 

(A) 
$$CH = CH_2$$

4. In which of the following reactions markovnikov rule of addition reaction is followed

$$\begin{array}{c|c} CH_3CH_3 \\ | & | \\ (A) CH_3 - C = C - CH_3 + HBr \longrightarrow \end{array}$$

$$\begin{array}{c|c}
CH_3 \\
| \\
(B) CH_3 - CH = C - CH_3 + Br_2 \xrightarrow{CCI_4}
\end{array}$$

5. The correct order of rate of following reactions is

$$CH_3 - C \equiv C - CH_3 \xrightarrow{\text{HBr (1 eq.)}} (X) \xrightarrow{\text{HBr (1 eq.)}} (Y)$$

$$H-C \equiv C-H \xrightarrow{HBr (1 \text{ eq.})} (Z)$$

$$CH_{3} - CH = CH - CH_{3} \xrightarrow{\text{HBr (1 eq.)}} (W)$$

$$(A) \ r_{4} > r_{2} > r_{1} > r_{3} \qquad (B) \ r_{1} > r_{2} > r_{3} > r_{4} \qquad (C) \ r_{4} > r_{3} > r_{2} > r_{1} \qquad (D) \ r_{3} > r_{4} > r_{2} > r_{4}$$

(A) 
$$r_4 > r_2 > r_1 > r_3$$

(B) 
$$r_1 > r_2 > r_3 > r$$

(C) 
$$r_4 > r_3 > r_2 > r_1$$

(D) 
$$r_3 > r_4 > r_2 > r_4$$

6. Which of the following shows least reactivity towards bromination?

$$\begin{array}{l} \text{(B) CH}_2\text{=}\text{CH}_2 \\ \text{(D) CH}_3\text{-}\text{CH=}\text{CH}_2 \end{array}$$

 $CH_3CH_2CH_2C \equiv CH \xrightarrow{BH_3, THF} \xrightarrow{H_2O_2/OH^-} 'X'$ 7. Identify the product 'X':

$$\begin{array}{c} \text{(B) } \mathsf{CH_3CH_2CH_2CCH_2} \\ \parallel \\ \mathsf{O} \end{array}$$

(C) 
$$CH_3CH_2CH_2CHO$$

8. The rate of acid catalysed hydration of II is much faster than that of I.

$$CH = CH - CH_3$$

$$CH_3$$
 $C = CH_2$ 
(II)

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(C)

6.

(C)

2.

7.

(A)

# **Hints & Solutions**

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As in previous question.

- 3. Addition of water by oxymercuration reduction without rearrangement.
- 4. Markovnikov addition is observed for unsymmetrical alkene and unsymmetrical reagent

- 5. Alkenes are more reactive than alkynes more branched alkynes are more reactive.
- 6. sp Hybridisation of alkyne and also intermediate form is less stable.

7. 
$$CH_3 - CH_2 - CH_2 - C \equiv CH \xrightarrow{(1)BH_3 - THF} CH_3 - CH_2 - CH_2 - CH = CH$$

Alkene II gives a tertiary carbocation intermediate in rate-determining step.

$$CH_3 \longrightarrow CH_3$$

$$CH_3 \longrightarrow CH_3$$

$$CH_3 \longrightarrow CH_3$$

3º carbocation